



US009045426B2

(12) **United States Patent**  
**Arimoto et al.**

(10) **Patent No.:** **US 9,045,426 B2**  
(45) **Date of Patent:** **Jun. 2, 2015**

(54) **METHOD FOR PRODUCING  
1,2-DIHYDROPYRIDINE-2-ONE COMPOUND**

(71) Applicant: **EISAI R&D MANAGEMENT CO.,  
LTD.**, Tokyo (JP)

(72) Inventors: **Itaru Arimoto**, Tokyo (JP); **Satoshi  
Nagato**, Tokyo (JP); **Yukiko Sugaya**,  
Tsukuba (JP); **Yoshio Urawa**, Kamisu  
(JP); **Koichi Ito**, Tsukuba (JP); **Hiroyuki**  
**Naka**, Kamisu (JP); **Takao Omae**,  
Tsukuba (JP); **Akio Kayano**, Kamisu  
(JP); **Katsutoshi Nishiura**, Kamisu (JP)

(73) Assignee: **EISAI R&D MANAGEMENT CO.,  
LTD.**, Tokyo (JP)

(\*) Notice: Subject to any disclaimer, the term of this  
patent is extended or adjusted under 35  
U.S.C. 154(b) by 0 days.

EP	1300396	A1	4/2003
EP	2053041	A2	4/2009
EP	2177520	A1	4/2010
IN	238440		2/2010
JP	52-83761	A	7/1977
JP	55-31072	A	3/1980
JP	2005-515995	A	6/2005
WO	WO 94/25469	A1	11/1994
WO	WO 95/01357	A1	1/1995
WO	WO 96/10023	A1	4/1996
WO	WO 96/33974	A1	10/1996
WO	WO 97/18163	A1	5/1997
WO	WO 97/28135	A1	8/1997
WO	WO 97/34878	A1	9/1997
WO	WO 97/43276	A1	11/1997
WO	WO 98/38173	A1	9/1998
WO	WO 98/50384	A1	11/1998
WO	WO 98/55480	A1	12/1998
WO	WO 99/31062	A1	6/1999
WO	WO 00/01376	A2	1/2000
WO	WO 00/07988	A1	2/2000
WO	WO 01/96308	A1	12/2001
WO	WO 03/047577	A2	6/2003

(21) Appl. No.: **14/165,365**

(22) Filed: **Jan. 27, 2014**

(65) **Prior Publication Data**

US 2014/0142315 A1 May 22, 2014

#### Related U.S. Application Data

(60) Division of application No. 12/870,507, filed on Aug.  
27, 2010, now Pat. No. 8,772,497, which is a division  
of application No. 11/649,299, filed on Jan. 4, 2007,  
now Pat. No. 8,304,548, which is a  
continuation-in-part of application No.  
PCT/JP2005/012364, filed on Jul. 5, 2005, and a  
continuation-in-part of application No.  
PCT/JP2005/012388, filed on Jul. 5, 2005.

(30) **Foreign Application Priority Data**

Jul. 6, 2004 (JP) ..... 2004-198709

(51) **Int. Cl.**  
**C07D 401/04** (2006.01)  
**C07D 213/64** (2006.01)

(52) **U.S. Cl.**  
CPC ..... **C07D 213/64** (2013.01)

(58) **Field of Classification Search**  
None  
See application file for complete search history.

(56) **References Cited**

#### U.S. PATENT DOCUMENTS

6,133,255 A 10/2000 Harrison et al.  
2004/0023973 A1 2/2004 Nagato et al.

#### FOREIGN PATENT DOCUMENTS

CN 1436172 A 8/2003  
DE 19643037 A1 4/1998  
EP 0802195 A2 10/1997

#### OTHER PUBLICATIONS

Vietnamese Notice of Allowance (including English translation  
thereof), dated Apr. 28, 2014, issued in Vietnamese Patent Applica-  
tion No. 1-2007-00266.

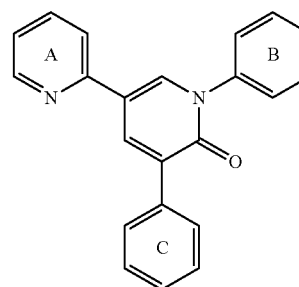
(Continued)

*Primary Examiner* — Zinna Northington Davis

(74) *Attorney, Agent, or Firm* — Birch, Stewart, Kolasch &  
Birch, LLP

(57) **ABSTRACT**

The present inventions provide a method for commercially  
producing a 1,2 -dihydropyridine-2-one compound repre-  
sented by the following formula (III-a)



(III-a)

wherein the ring A represents an optionally substituted 2-py-  
ridyl group, the ring B represents an optionally substituted  
phenyl group, and the ring C represents an optionally substi-  
tuted phenyl group. Further, the invention provides crystals of  
3-(2-cyanophenyl)-5-(2-pyridyl)-1-phenyl-1,2 -dihydropyri-  
din-2-one and production processes therefore.

**20 Claims, 12 Drawing Sheets**

(56)

**References Cited****OTHER PUBLICATIONS**

Jordanian Office Action and English translation thereof, dated May 6, 2014, for Jordanian Application No. 97/2005.

Notice of Allowance for U.S. Appl. No. 12/870,507 dated Jan. 14, 2014.

Supplemental Notice of Allowability issued Apr. 11, 2014, in U.S. Appl. No. 12/870,507.

Shelving Notification issued Jun. 10, 2014, in Brazilian Patent Application No. BR1220130308067, with English translation.

Notice of Allowance for co-pending U.S. Appl. No. 12/870,507, dated Mar. 11, 2014.

Notice of Grant issued Feb. 24, 2014 for corresponding Chinese application No. 201110261753.4 with English translation.

Arias et al., "β-Amyloid peptide fragment 25-35 potentiates the calcium-dependent release of excitatory amino acids from depolarized hippocampal slices," *Journal of Neuroscience Research*, vol. 41, 1995, pp. 561-566.

Australian Amendment, filed Dec. 18, 2006, for Australian Patent Application No. 2005258385.

Australian Examination Report issued in Australian Application No. 2005258378 on Jul. 26, 2010.

Australian Notice of Acceptance, dated Oct. 27, 2009, for Australian Patent Application No. 2005258385.

Australian Office Action, dated Jul. 10, 2009, for Australian Patent Application No. 2005258385.

Australian Office Action, dated Nov. 3, 2008, for Australian Patent Application No. 2005258385.

Australian Response to Office Action, filed Jun. 10, 2009, for Australian Patent Application No. 2005258385.

Australian Response to Office Action, filed Sep. 14, 2009, for Australian Patent Application No. 2005258385.

Australian Response, dated Mar. 29, 2011, to Australian Office Action issued in Australian Application No. 2005258378 on Jul. 26, 2010.

Bangladeshi Examination Report, dated Feb. 28, 2007, for Bangladeshi Patent Application No. 147/2005.

Bangladeshi Notice of Allowance, dated May 8, 2007, for Bangladeshi Patent Application No. 147/2005.

Bangladeshi Notice of Allowance, dated May 8, 2007, for Bangladeshi Patent Application No. 47/2007.

Bangladeshi Response to Examination Report, filed Mar. 7, 2007, for Bangladeshi Patent Application No. 147/2005.

Brazilian Communication dated Aug. 2, 2012 from counterpart Brazilian Patent Application No. PI0510254-5, with English translation.

Brazilian Office Action for Brazilian Application No. PI0510254-5 dated Apr. 17, 2013, with English translation.

Brazilian Office Action for Brazilian Application No. PI0510254-5 dated Jun. 5, 2012, with English translation.

Brazilian Submission Document(s) Before the Patent Office, dated May 17, 2013, in Brazilian Application No. PI0510254-5, with English translation.

Brusa et al., "Early-Onset Epilepsy and Postnatal Lethality Associated with an Editing-Deficient GluR-B Allele in Mice," *Science*, vol. 270, Dec. 1995, pp. 1677-1680.

Canadian Amendment After Allowance, filed Aug. 25, 2009, for Canadian Patent Application No. 2,561,829.

Canadian Notice of Allowance, dated Apr. 14, 2009, for Canadian Patent Application No. 2,561,829.

Canadian Office Action issued in Canadian Application No. 2,570,177 on Feb. 3, 2011.

Canadian Office Action issued in Canadian Application No. 2,570,177 on Jun. 9, 2011.

Canadian Response to Canadian Office Action issued in Canadian Application No. 2,570,177 on Feb. 3, 2011.

Canadian Response, dated Jul. 6, 2011, to Canadian Office Action issued in Canadian Application No. 2,570,177 on Jun. 9, 2011.

Castillo et al., "Progression of ischemic stroke and excitotoxic aminoacids," *The Lancet*, vol. 349, 1997, pp. 79-83.

Cha et al., "2,4,5-Trihydroxyphenylalanine (6-hydroxy-DOPA) displaces [3H]AMPA binding in rat striatum," *Neuroscience Letters*, vol. 132, 1991, pp. 55-58.

Chinese Amendment filed on Dec. 15, 2011, for Chinese Application No. 201010246226.1, with English translation (Jan. 1, 2012).

Chinese Notice of Allowance dated Jan. 13, 2012 for corresponding Chinese Patent Application No. 200580011565.9, with English translation.

Chinese Notice of Allowance issued on Feb. 9, 2012, in counterpart Chinese Application No. 200910138575.9.

Chinese Notification to Complete Registration Formalities and Notification to Grant Patent Right for Invention, dated Apr. 5, 2012, for Chinese Application No. 201010246226.1, with English translation.

Chinese Office Action dated Jun. 4, 2013 issued in counterpart Chinese Patent Application No. 201110261753.4.

Chinese Office Action issued in Chinese Application No. 200580023071.2 on Aug. 7, 2009, with English translation.

Chinese Office Action issued in Chinese Application No. 200580023071.2 on Dec. 6, 2010, with English translation.

Chinese Office Action issued in Chinese Application No. 200580023071.2 on Feb. 27, 2009, with English translation.

Chinese Office Action issued in Chinese Application No. 200580023071.2 on May 11, 2010.

Chinese Office Action issued in Chinese Application No. 200580023071.2 on Nov. 20, 2009.

Chinese Office Action, dated Apr. 28, 2010, for Chinese Patent Application No. 200580011565.9, with English translation.

Chinese Office Action, dated Dec. 23, 2010, for Chinese Patent Application No. 200580011565.9, with English translation.

Chinese Office Action, dated Dec. 31, 2011, issued in corresponding Chinese Application No. 201010246226.1, with English translation.

Chinese Office Action, dated Jun. 9, 2011, for Chinese Patent Application No. 201010246226.1, with English translation.

Chinese Office Action, dated May 6, 2010, for Chinese Patent Application No. 200910138575.9, with English translation.

Chinese Office Action, dated Nov. 21, 2008, for Chinese Patent Application No. 200580011565.9, with English translation.

Chinese Office Action, dated Oct. 19, 2011, for Chinese Patent Application No. 200910138575.9, with English translation.

Chinese Rejection Decision issued in Chinese Application No. 200580023071.2 on May 25, 2011, with English translation.

Chinese Request for Re-examination of Chinese Office Action issued in Chinese Application No. 200580023071.2 on May 25, 2011, with English translation.

Chinese Response to Chinese Office Action issued in Chinese Application No. 200580023071.2 on Aug. 7, 2009, with English translation.

Chinese Response to Chinese Office Action issued in Chinese Application No. 200580023071.2 on Dec. 6, 2010, with English translation.

Chinese Response to Chinese Office Action issued in Chinese Application No. 200580023071.2 on Feb. 27, 2009, with English translation.

Chinese Response to Chinese Office Action issued in Chinese Application No. 200580023071.2 on May 11, 2010, with English translation.

Chinese Response to Chinese Office Action issued in Chinese Application No. 200580023071.2 on Nov. 12, 2009, with English translation.

Chinese Response to Office Action, filed Aug. 4, 2010, for Chinese Patent Application No. 200910138575.9, with English translation.

Chinese Response to Office Action, filed Feb. 21, 2011, for Chinese Patent Application No. 200580011565.9, with English translation.

Chinese Response to Office Action, filed Jul. 12, 2010, for Chinese Patent Application No. 200580011565.9, with English translation.

Chinese Response to Office Action, filed May 7, 2009, for Chinese Patent Application No. 200580011565.9, with English translation.

Chinese Response to Office Action, filed Nov. 16, 2011, for Chinese Patent Application No. 200910138575.9, with English translation.

Chinese Response to Office Action, filed Sep. 20, 2011, for Chinese Patent Application No. 201010246226.1, with English translation.

Chinese Response to the Chinese Office Action filed on Feb. 22, 2012, for Application No. 201010246226.1, with English translation.

(56)

**References Cited****OTHER PUBLICATIONS**

Chinese Submission of Document(s) Before the Patent Office, dated Sep. 27, 2013, in counterpart Chinese Application No. 201110261753.4.

Chinese Supplemental Response with Experimental Data to Chinese Office Action issued in Chinese Application No. 200580023071.2 on May 11, 2010, with English translation.

Chinese Voluntary Argument and Amendment, dated Dec. 15, 2011, for Chinese Patent Application No. 201010246226.1, with English translation.

Dreyer et al., "The Coat Protein gp120 of HIV-1 Inhibits Astrocyte Uptake of Excitatory Amino Acids via Macrophage Arachidonic Acid," *European Journal of Neuroscience*, vol. 7, 1995, pp. 2502-2507.

Earl et al., "The preparation of 2(1H)-pyridinones and 2,3-dihydro-5(1H)-indolizinones via transition metal mediated cocyclization of alkynes and isocyanates. A novel construction of the antitumor agent. camptothecin," *Journal of Organic Chemistry*, vol. 49, 1984, pp. 4786-4800.

Egyptian Office Action (date of dispatch: unknown) for Egyptian Application No. 1270/2006 in review of a Response dated Mar. 3, 2010.

Egyptian Office Action, dated Dec. 14, 2010, for Egyptian Patent Application No. 1270/2006, with English translation.

Egyptian Office Action, dated Dec. 15, 2009, for Egyptian Patent Application No. 1270/2006, with English translation.

Egyptian Office Action, dated Jul. 30, 2007, for Egyptian Patent Application No. 1270/2006, with English translation.

Egyptian Response to Office Action, filed Mar. 3, 2010, for Egyptian Patent Application No. 1270/2006, with English translation.

Egyptian Response to Office Action, filed on Aug. 25, 2007, for Egyptian Patent Application No. 1270/2006, with English translation.

El-Kholy, "Pyrone series. Part IV. 5-Aryloxy-2-pyrones. The corresponding 2-thiopyrones, pyridones, 1-hydroxy- and 1-amino-2-pyridones, and related 4,5,6-triphenyl-2-pyrones," *J. Chem. Soc.*, 1961, pp. 4490-4498.

English translation of Applicant Response to Israeli Office Action issued in Israeli Application No. 180151 on Oct. 24, 2010.

European Communication for European Application No. 05758232.2, dated May 14, 2012.

European Examination Report issued in European Application No. 05758231.4 on Aug. 11, 2009.

European Notice of Allowance dated Jan. 7, 2013, for European Application No. 05758232.2.

European Notice of Allowance dated Nov. 9, 2012 for European Application No. 05758232.2.

European Response, dated Dec. 21, 2009, to European Office Action issued in European Application No. 05 758 231.4 on Aug. 11, 2009.

European Submission Documents dated Aug. 9, 2012 Before the European Patent Office for counterpart European Application No. 05 758 232.2.

European Supplemental Response, dated Oct. 27, 2010, to European Office Action issued in European Application No. 05 758 231.4 on Aug. 11, 2009.

Extended European Search Report dated Apr. 25, 2012 issued in counterpart European Application No. 05758232.2.

Extended European Search Report dated Jul. 3, 2013 issued in counterpart EP Patent Application No. 13152837.4.

Extended European Search Report issued in European Application No. 05758231.4 on May 12, 2009.

Filipino Notice of Allowability, dated Feb. 26, 2010, for Filipino Patent Application No. 1-2006-501797.

Filipino Office Action, dated Dec. 17, 2009, for Filipino Patent Application No. 1-2006-501797.

Filipino Response to Office Action, dated Jan. 21, 2010, for Filipino Patent Application No. 1-2006-501797.

Fujita et al., "Studies on 1-Alkyl-2(1H)-pyridone Derivatives XXXII. The Friedel-Crafts Reaction of 1-Alkyl-2(1H)-pyridone Derivatives with Acid Anhydride," *Yagugaku Zasshi*, vol. 110, No. 6, 1990, pp. 449-452.

Gahring et al., "Autoantibodies to glutamate receptor subunit GluR2 in nonfamilial olivopontocerebellar degeneration," *Neurology*, vol. 48, 1997, pp. 494-500.

Gissot et al., "A Functionalized, Deep Cavitand Catalyzes the Aminolysis of a Choline Derivative" *Journal of the American Chemical Society*, vol. 126, No. 24, 2004, pp. 7424-7425, XP002522905.

Gray et al., "A role for AMPA/kainate receptors in conditioned place preference induced by diazepam in the rat," *Neuroscience Letters*, vol. 268, 1999, pp. 127-130.

Haleblian et al., "Pharmaceutical Applications of Polymorphism," *Journal of Pharmaceutical Sciences*, vol. 58, No. 8, Aug. 1969, pp. 911-929.

Hardin-Pouzet et al., "Glutamate metabolism is down-regulated in astrocytes during experimental allergic encephalomy," *GLIA*, vol. 20, 1997, pp. 79-85.

Indian Examination Report, dated Nov. 13, 2009, for Indian Patent Application No. 533/CHENP/2007.

Indian First Examination Report, dated Dec. 15, 2008, for Indian Patent Application No. 533/CHENP/2007.

Indian Notice of Hearing issued in Indian Application No. 7715/DELNO/2006 on Sep. 5, 2011.

Indian Office Action issued in Indian Application No. 7715/DELNP/2006 on Dec. 14, 2009 (stamped Dec. 15, 2009).

Indian Office Action issued in Indian Application No. 7716/DELNP/2006 on Sep. 8, 2010.

Indian Response to Examination Report, filed Dec. 8, 2009, for Indian Patent Application No. 533/CHENP/2007.

Indian Response to First Examination Report, filed Apr. 13, 2009, for Indian Patent Application No. 533/CHENP/2007.

Indian Response to Indian Notice of Hearing issued in Indian Application No. 7715/DELNP/2006 on Sep. 5, 2011.

Indian Response to Indian Office Action issued in Indian Application No. 7715/DELNP/2006 on Aug. 8, 2010.

Indian Response to Indian Office Action issued in Indian Application No. 7715/DELNP/2006 on Dec. 15, 2009.

International Preliminary Report on Patentability and Written Opinion of the International Searching Authority, dated Jan. 16, 2007, for International Patent Application No. PCT/JP2005/012388.

International Search Report, dated Sep. 20, 2005, for International Patent Application No. PCT/JP2005/012388.

Israeli Notice Prior to Allowance, dated Jul. 27, 2011, for Israeli Patent Application No. 180024, with English translation.

Israeli Notice Prior to Examination, dated Mar. 3, 2009, for Israeli Patent Application No. 180024, with English translation.

Israeli Office Action issued in Israeli Application No. 180151 on Oct. 24, 2010, with English translation.

Israeli Office Action, dated Jul. 12, 2010, for Israeli Patent Application No. 180024, with English translation.

Israeli Response to Notice Prior to Examination, filed Jun. 25, 2009, for Israeli Patent Application No. 180024, with English translation.

Israeli Response to Office Action, filed Sep. 6, 2010, for Israeli Patent Application No. 180024, with English translation.

Israeli Response, dated Jan. 25, 2011, to Israeli Office Action, dated Oct. 24, 2010, issued in Israeli Application No. 180151, with English translation.

Japanese Amendment and Argument filed on Oct. 17, 2013 in counterpart Japanese Patent Application No. 2011-224926, with English translation.

Japanese Amendment and Argument filed on Sep. 6, 2013 for counterpart Japanese Patent Application No. 2011-224926, with English translation.

Japanese Amendment and Argument, dated Apr. 17, 2012, for Japanese Application No. 2006-528897, with English translation.

Japanese Amendment, dated May 30, 2012, for Japanese Application No. 2006-528897, with English translation.

Japanese Certificate of Experiment Results, dated Apr. 16, 2012, for Japanese Application No. 2006-528897, with English translation.

Japanese Decision of Appeal, dated Jun. 22, 2012, for Japanese Application No. 2006-528897, with English translation.

(56)

**References Cited****OTHER PUBLICATIONS**

Japanese Decision of Rejection, dated Sep. 17, 2010, for Japanese Application No. 2006-568897, with English translation.

Japanese Demand for Appeal, dated Dec. 16, 2010, of Decision of Rejection, dated Sep. 13, 2010, for Japanese Application No. 2006-528897, with English translation.

Japanese Notice of Decision to Grant a Patent, dated Nov. 5, 2013, in counterpart Japanese Application No. P2011-224926, with English translation.

Japanese Notification of Reason for Rejection mailed Jun. 11, 2010, for corresponding Japanese Application No. 2006-528897, with English translation.

Japanese Notification of Reason for Rejection, dated Feb. 17, 2012, for Japanese Application No. 2006-528897 (Appeal No. 2010-28433), with English translation.

Japanese Notification of Reason for Rejection, dated May 29, 2012, for Japanese Application No. 2006-528897, with English translation.

Japanese Office Action for Application No. 2007-009518 dated Aug. 16, 2011, with English translation thereof and English translation of Claims for Japanese Application No. 2007-009518.

Japanese Office Action, dated Jul. 9, 2013, issued in counterpart Japanese Application No. P2011-224926, with English translation.

Japanese Office Action, dated May 8, 2007, for Japanese Application No. 2006-528904, with English translation.

Japanese Office Action, dated Nov. 21, 2006, for Japanese Application No. 2006-528904, with English translation.

Japanese Office Action, dated Oct. 1, 2013, for counterpart Japanese Application No. 2011-224926, with English translation.

Japanese Re-Corrected Comparative Experiment Report, dated Apr. 16, 2012, for Japanese Application No. 2006-528897, with English translation.

Japanese Response, dated Jul. 30, 2010, to Japanese Office Action issued in Japanese Application No. 2006-528897 on Jun. 2, 2010 (mailed Jun. 11, 2010), with English translation.

Japanese Written Amendment, filed Jan. 18, 2007, for Japanese Application No. 2006-528904, with English translation.

Japanese Written Amendment, filed Oct. 23, 2006, for Japanese Patent Application No. 2006-528904, with English translation.

Japanese Written Argument, filed Jan. 18, 2007, for Japanese Application No. 2006-528904, with English translation.

Japanese Written Statement, filed Nov. 2, 2006, for Japanese Application No. 2006-528904, with English translation.

Japanese Written Statement, filed Oct. 23, 2006, for Japanese Application No. 2006-528904, with English translation.

Kaczmarek et al., "Bulletin of the Polish Academy of Sciences", Beilstein Institute for Organic Chemistry, vol. 32, No. 3-6, Jan. 2009, pp. 105-114, XP-002522906.

Korean Amendment, filed Jun. 2, 2009, for Korean Patent Application No. 10-2006-7020395, with English translation.

Korean Amendment, filed Oct. 11, 2006, for Korean Patent Application No. 10-2006-7020395, with English translation.

Korean Argument and Amended Claims, dated Mar. 27, 2012, for Korean Application No. 10-2006-7027447, with English translation.

Korean Argument Brief, filed Jun. 2, 2009, for Korean Patent Application No. 10-2006-7020395, with English translation.

Korean Notice of Allowance, dated Jul. 31, 2012, for Korean Application No. 10-2006-7027447, with English translation.

Korean Notice of Decision for Patent, dated Oct. 29, 2009, for Korean Patent Application No. 10-2006-7020395, with English translation.

Korean Notice to Submit a Response, dated Jan. 27, 2012, for Korean Application No. 10-2006-7027447, with English translation.

Korean Office Action, dated May 11, 2011, for Korean Application No. 10-2006-7027447, with English translation.

Korean Office Action, dated Sep. 3, 2007, for Korean Patent Application No. 10-2006-7020395, with English translation.

Korean Response, dated Jul. 11, 2011, to Korean Office Action issued in Korean Application No. 5-2006-031192-0 on May 11, 2011, with English translation.

Kotlinska et al., "The Putative AMPA Receptor Antagonist, LY326325, Produces Anxiolytic-Like Effects Without Altering

Locomotor Activity in Rats," *Pharmacology Biochemistry and Behavior*, vol. 60, No. 1, 1998, pp. 119-124.

Lipton, "Memantine prevents HIV coat protein-induced neuronal injury in vitro," *Neurology*, vol. 42, 1992, pp. 1403-1405.

Macau Request to Record the Extended Patent Application, dated May 29, 2012, for Macau Application No. J/000789(395).

Malaysian Amendment, filed Dec. 21, 2009, for Malaysian Patent Application No. PI20052537.

Malaysian Notice of Allowance dated Jun. 7, 2013 issued in counterpart Malaysian Patent Application No. PI20052537.

Meldrum, "Amino Acids as Dietary Excitotoxins: A Contribution to Understanding Neurodegenerative Disorders," *Brain Research Reviews*, vol. 18, 1993, pp. 293-314.

Mexican Notice of Allowance, dated Mar. 9, 2012, for Mexican Patent Application No. PA/a/2006/014458, with English translation.

Mirza et al., "Influence of Solvents on the Variety of Crystalline Forms of Erythromycin," *AAPS PharmSci*, vol. 5, No. 2, Article 12, Apr. 14, 2003, pp. 1-9.

Muller et al., "gp120 of HIV-1 induces apoptosis in rat cortical cell cultures: prevention by memantine," *European Journal of Pharmacology*, vol. 226, 1992, pp. 209-214.

Nadin et al., "Synthesis of Tricyclic Pyridones by Radical Cyclization," *Tetrahedron Letters*, vol. 40, No. 21, 1999, pp. 4073-4076.

Nelson et al., "Specific High Molecular weight mRNAs induced by associative learning in *Hermissenda*," *Proc. Natl. Acad. Sci.*, vol. 87, 1990, pp. 269-273.

New Zealand Examination Report and Notice of Acceptance, dated Oct. 7, 2009, for New Zealand Patent Application No. 550720.

New Zealand Examination Report, dated Apr. 30, 2009, for New Zealand Patent Application No. 550720.

New Zealand Examination Report, dated Sep. 9, 2009, for New Zealand Patent Application No. 550720.

New Zealand Response to Examination Report, filed Aug. 24, 2009, for New Zealand Patent Application No. 550720.

New Zealand Response to Examination Report, filed Sep. 15, 2009, for New Zealand Patent Application No. 550720.

Pakistani Notice of Acceptance, dated Nov. 5, 2010, for Pakistani Patent Application No. 618/2005.

Pakistani Office Action, dated Jun. 18, 2007, for Pakistani Patent Application No. 618/2005.

Pakistani Response to Office Action, filed Oct. 14, 2010, for Pakistani Patent Application No. 618/2005.

Partial European Search Report dated Mar. 5, 2013 for European Application No. 13152837.4.

Pessolano et al., "Novel Nucleophilic Substitution of Alkyl Bromo-2(1H)-pyridones," *Journal of Heterocyclic Chemistry*, vol. 22, 1985, pp. 265-272, XP-002522914.

Rosenberg et al., "2,4,5-Trihydroxyphenylalanine in solution forms a non-N-methyl-D-aspartate glutamatergic agonist and neurotoxin," *Proc. Natl. Acad. Sci.*, vol. 88, 1991, pp. 4865-4869.

Russian Decision on Grant, dated Oct. 30, 2007, for Russian Patent Application No. 2007104342, with English translation.

Saudi Arabian Appeal before Committee, filed Jan. 26, 2010, for Saudi Arabian Patent Application No. 5260199, with English translation.

Saudi Arabian Examination Report, dated Aug. 26, 2009, for Saudi Arabian Patent Application No. 5260199, with English translation.

Saudi Arabian Examination Report, dated Dec. 7, 2009, for Saudi Arabian Patent Application No. 5260199, with English translation.

Saudi Arabian Examination Report, dated May 13, 2009, for Saudi Arabian Patent Application No. 5260199, with English translation.

Saudi Arabian Rejection from Committee, dated May 16, 2010, for Saudi Arabian Patent Application No. 5260199, with English translation.

Saudi Arabian Response to Examination Report, dated Jul. 27, 2009, for Saudi Arabian Patent Application No. 5260199, with English translation.

Saudi Arabian Response to Examination Report, dated Nov. 14, 2009, for Saudi Arabian Patent Application No. 5260199, with English translation.

Sheardown et al., "2,3-Dihydroxy-6-nitro-7-sulfamoyl-benzo(F)quinoxaline: a neuroprotectant for cerebral ischemia," *Science*, vol. 247, 1990, pp. 571-574.

(56)

**References Cited****OTHER PUBLICATIONS**

Sindou et al., "Prevention of HIV coat protein (gp120) toxicity in cortical cell cultures by riluzole," *Journal of Neurological Sciences*, vol. 126, 1994, pp. 133-137.

Smith et al., "L-DOPA increases nigral production of hydroxyl radicals in vivo: potential L-DOPA toxicity?", *NeuroReport*, vol. 5, 1994, pp. 1009-1011.

South African Amendment, filed Jan. 11, 2007, for South African Patent Application No. 2006/09274.

South African Notice of Acceptance, dated Apr. 24, 2008, for South African Patent Application No. 2006/09274.

Taiwanese Amendment, filed May 18, 2011, for Taiwanese Patent Application No. 094122578, with English translation.

Taiwanese Notice of Allowance, dated Jun. 2, 2011, for Taiwanese Patent Application No. 094122578, with English translation.

Taiwanese Office Action, dated Feb. 21, 2011, for Taiwanese Patent Application No. 094122578, with English translation.

Taiwanese Response, filed May 18, 2011, for Taiwanese Patent Application No. 094122578, with English translation.

Thai Amendment, filed Feb. 23, 2011, for Thai Patent Application No. 0501003082, with English translation.

Thathagar et al., "Supporting information for the article 'Copper-catalysed Suzuki Cross-coupling using Mixed Nanocluster Catalysts,' pp. S1-S4," *J Am Chem Soc*, Oct. 9, 2002, vol. 124, No. 40, pp. 11858-11859, [http://pubs.acs.org/doi/suppl/10.1021/ja027716%2B/suppl\\_file/ja027716%2B\\_s1.pdf](http://pubs.acs.org/doi/suppl/10.1021/ja027716%2B/suppl_file/ja027716%2B_s1.pdf).

Turski et al., "Relief of experimental spasticity and anxiolytic/anticonvulsant actions of the alpha-amino-3-hydroxy-5-methyl-4-

isoxazolepropionate antagonist 2,3-dihydroxy-6-nitro-7-sulfamoyl-benzo(F)quinoxaline," *J Pharmacol Exp Ther*, vol. 260, No. 2, 1992, pp. 742-747.

Ushijima et al., "Exposure to gp120 of HIV-1 Induces an Increased Release of Arachidonic Acid in Rat Primary Neuronal Cell Culture Followed by NMDA Receptor-mediated Neurotoxicity," *European Journal of Neuroscience*, vol. 7, 1995, pp. 1353-1359.

Vietnamese Response to Office Action, filed Jul. 10, 2008, for Vietnamese Patent Application No. 1-2007-00266, with English translation.

Vietnamese Office Action, dated Jun. 5, 2008, for Vietnamese Patent Application No. 1-2007-00266, with English translation.

Vietnamese Response, dated Oct. 5, 2012, to Vietnamese Official Notice for Vietnamese Application No. 1-2007-00266, with English translation.

Yoshiyama et al., "Effects of LY215490, a competitive alpha-amino-3-hydroxy-5-methylisoxazole-4-propionic acid (AMPA) receptor antagonist, on the micturition reflex in the rat," *J Pharmacol Exp Ther*, vol. 280, No. 2, 1997, pp. 894-904.

Notice of Allowance issued Sep. 15, 2014, in Jordan Patent Application No. 2005/97.

Office Action issued Oct. 8, 2014, in Indonesian Patent Application No. W-00 2007 00022, with English translation.

Voluntary Amendment filed Sep. 22, 2014, in Indonesian Patent Application No. W-00 2007 00022, with English translation.

First Examination Report issued Dec. 22, 2014, in Indian Patent Application No. 7210/CHENP/2009.

Reply filed Jan. 8, 2015, in response to the Office Action issued Oct. 8, 2014, in Indonesian Patent Application No. W-00 2007 00022, with English translation.

Notice of Allowance for Indonesian Patent Application No. W-00200700022 dated Feb. 20, 2015, with English translation.

FIG. 1

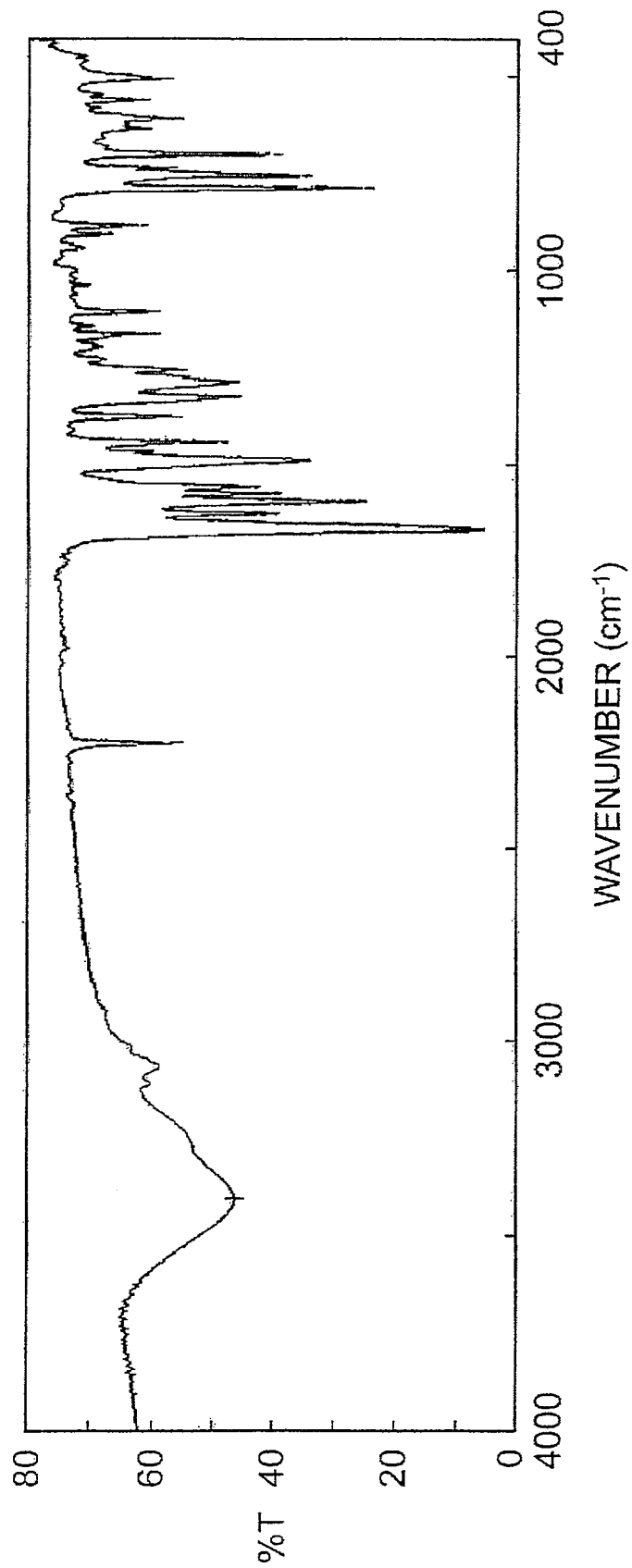


FIG. 2

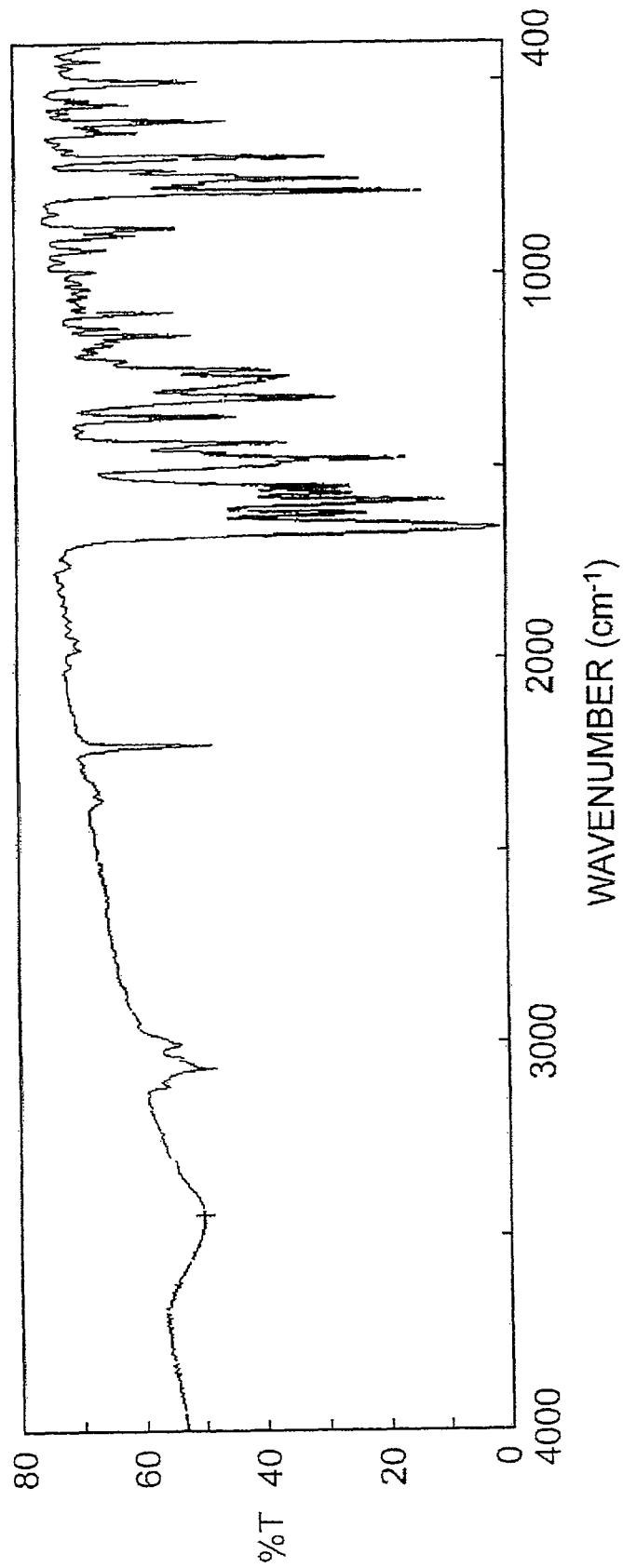


FIG. 3

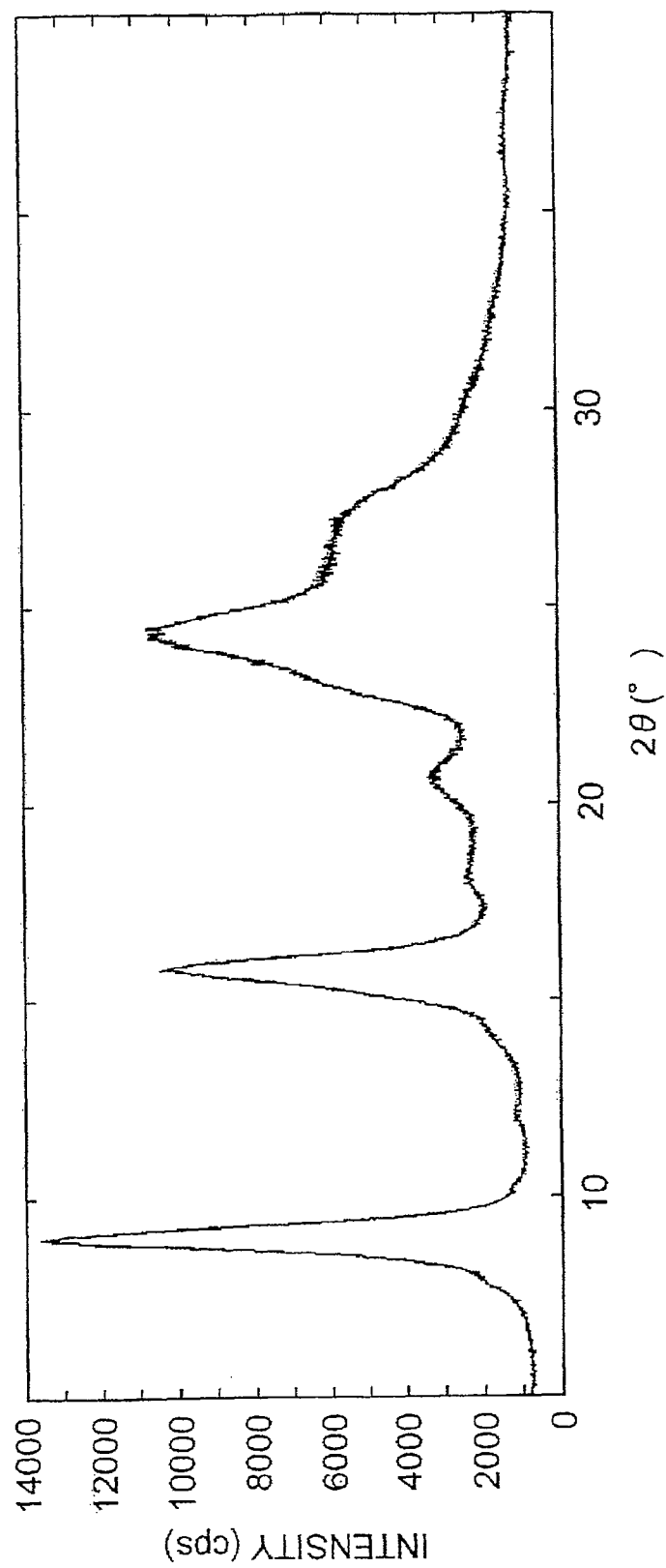




FIG. 4

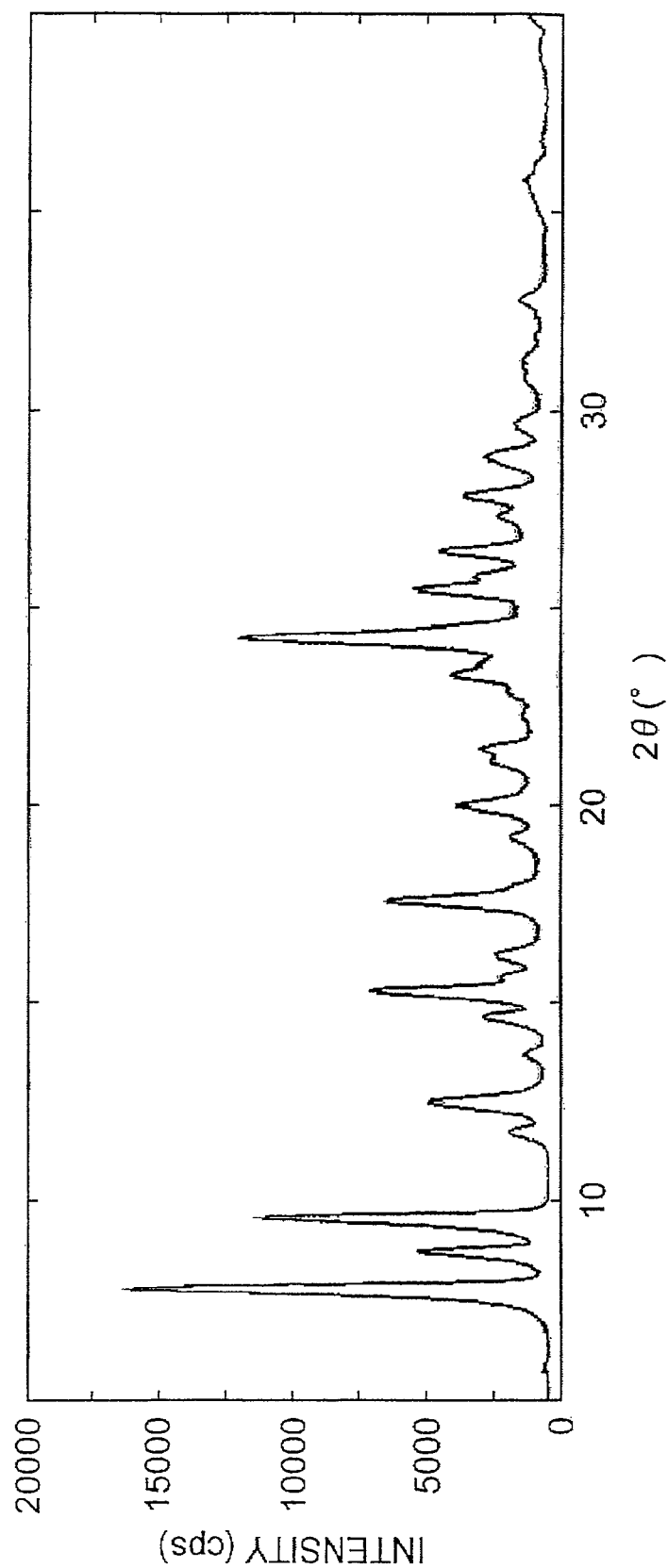


FIG. 5

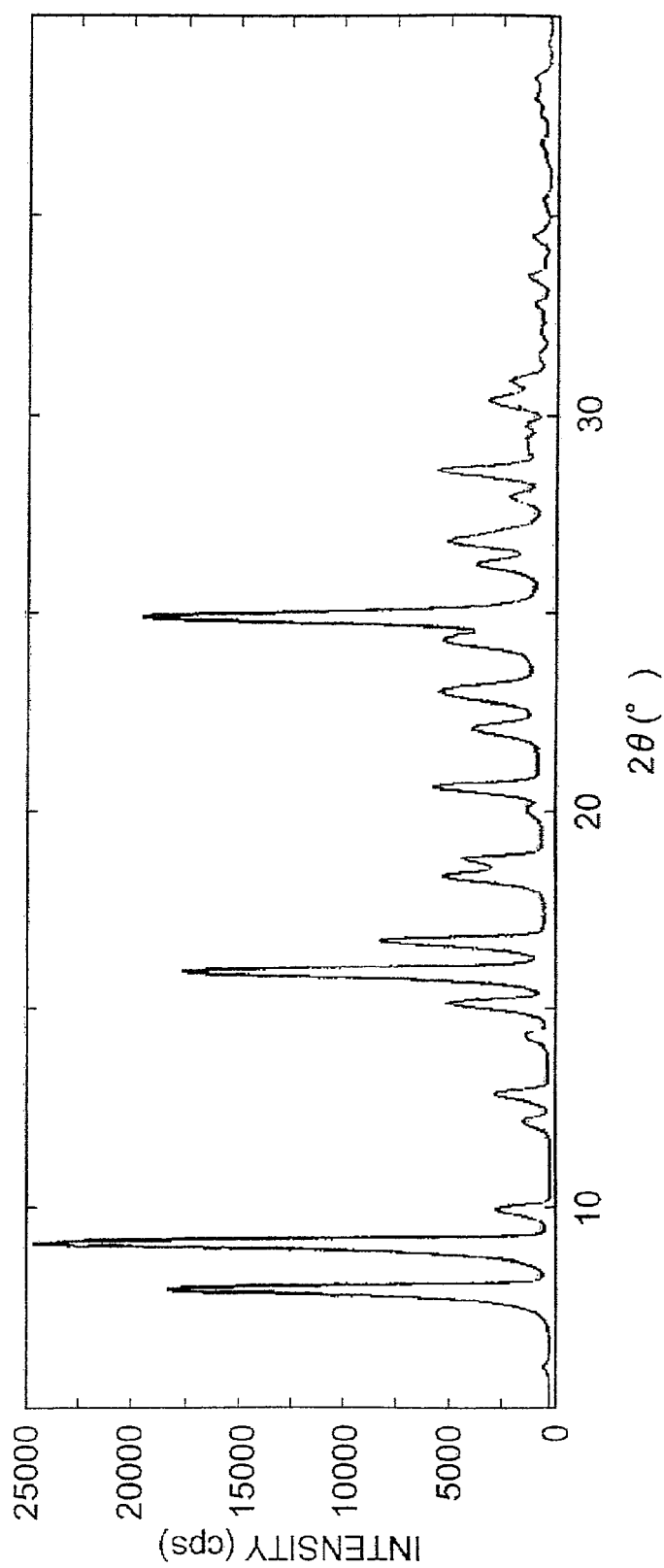


FIG. 6

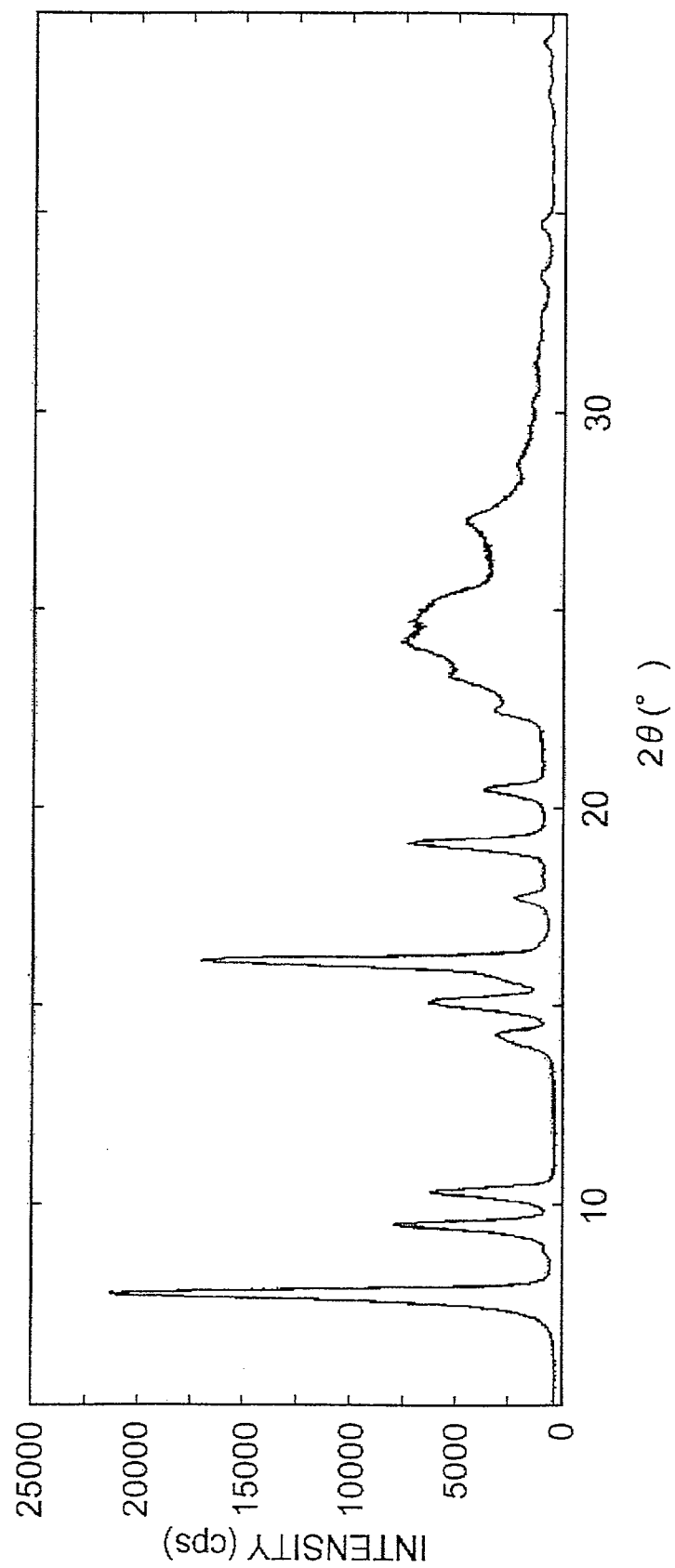


FIG. 7

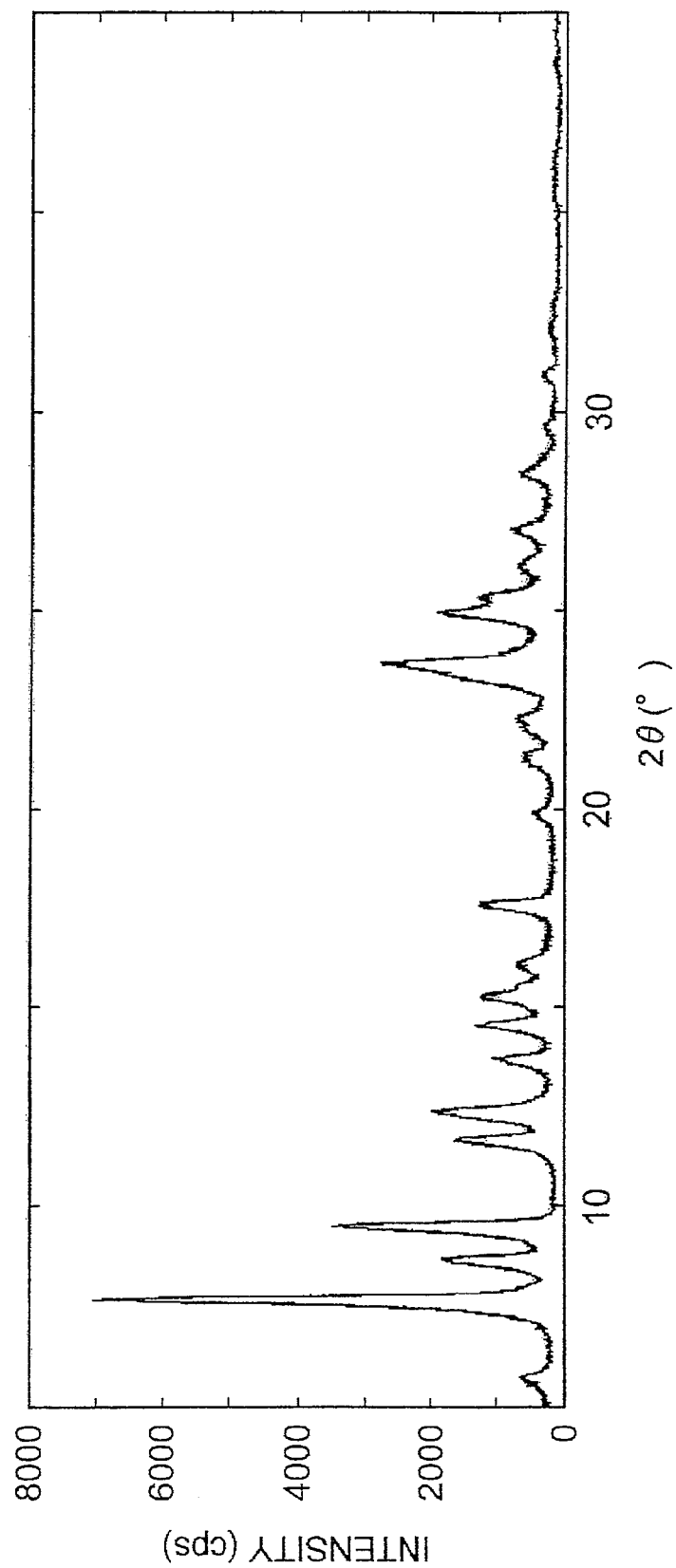


FIG. 8

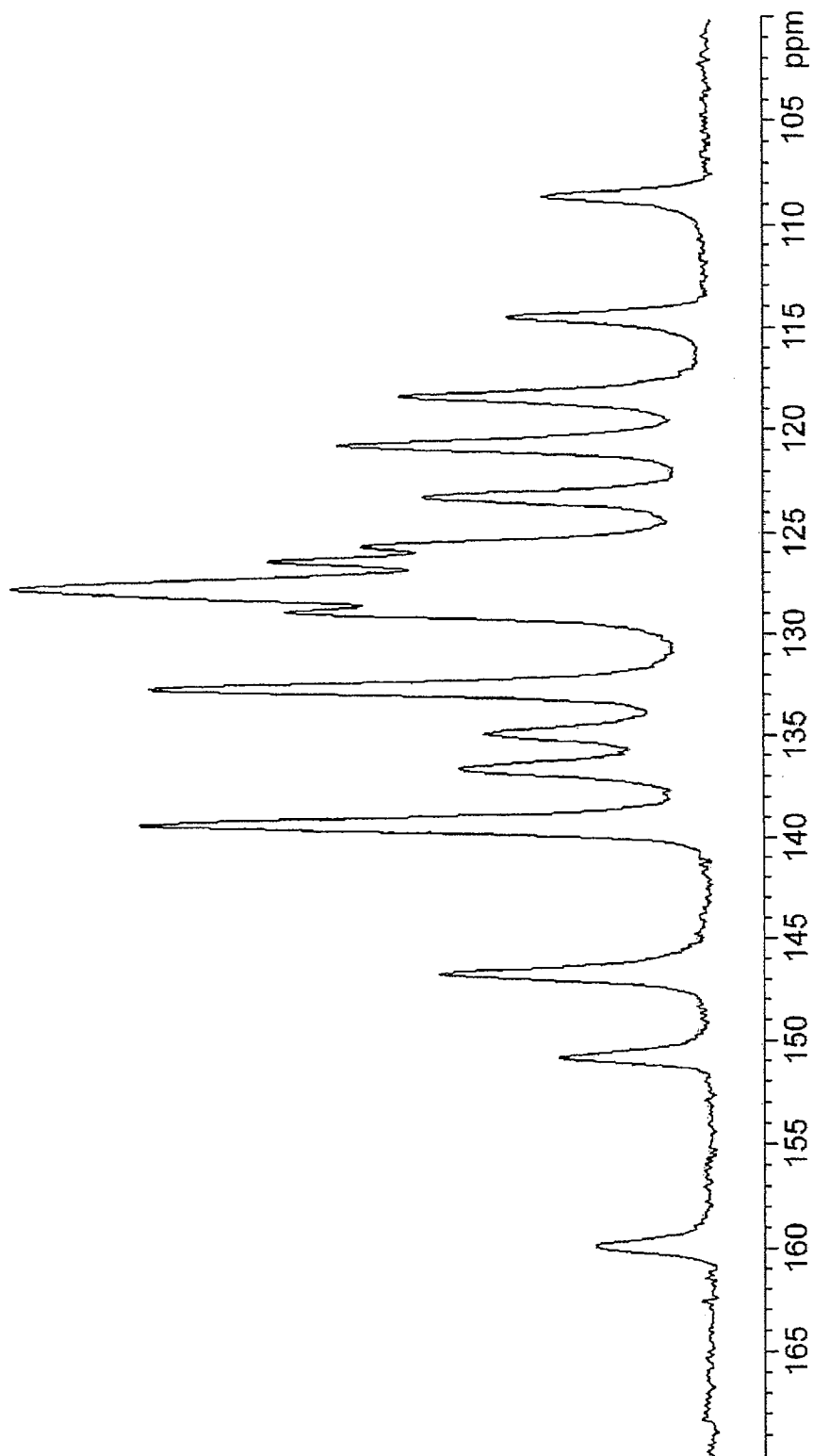


FIG. 9

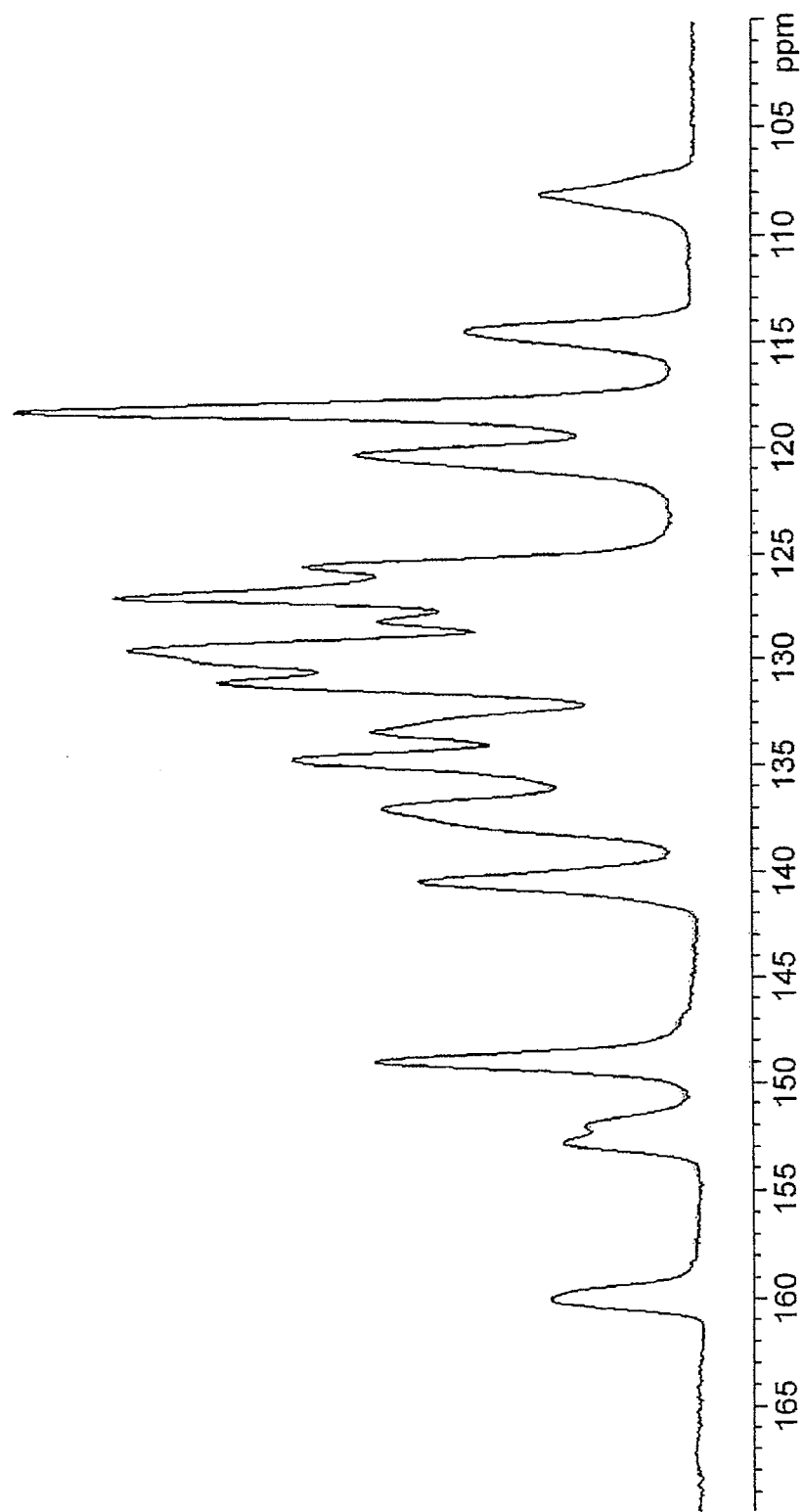


FIG. 10

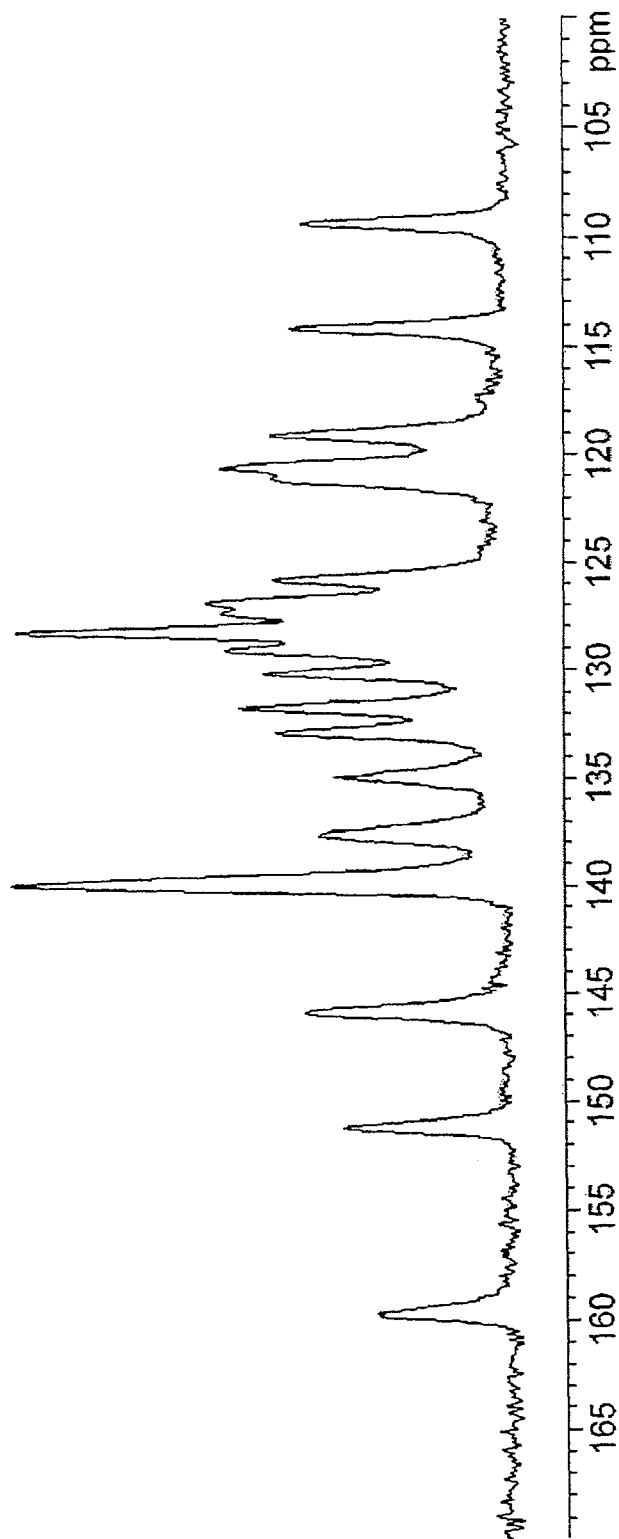


FIG. 11

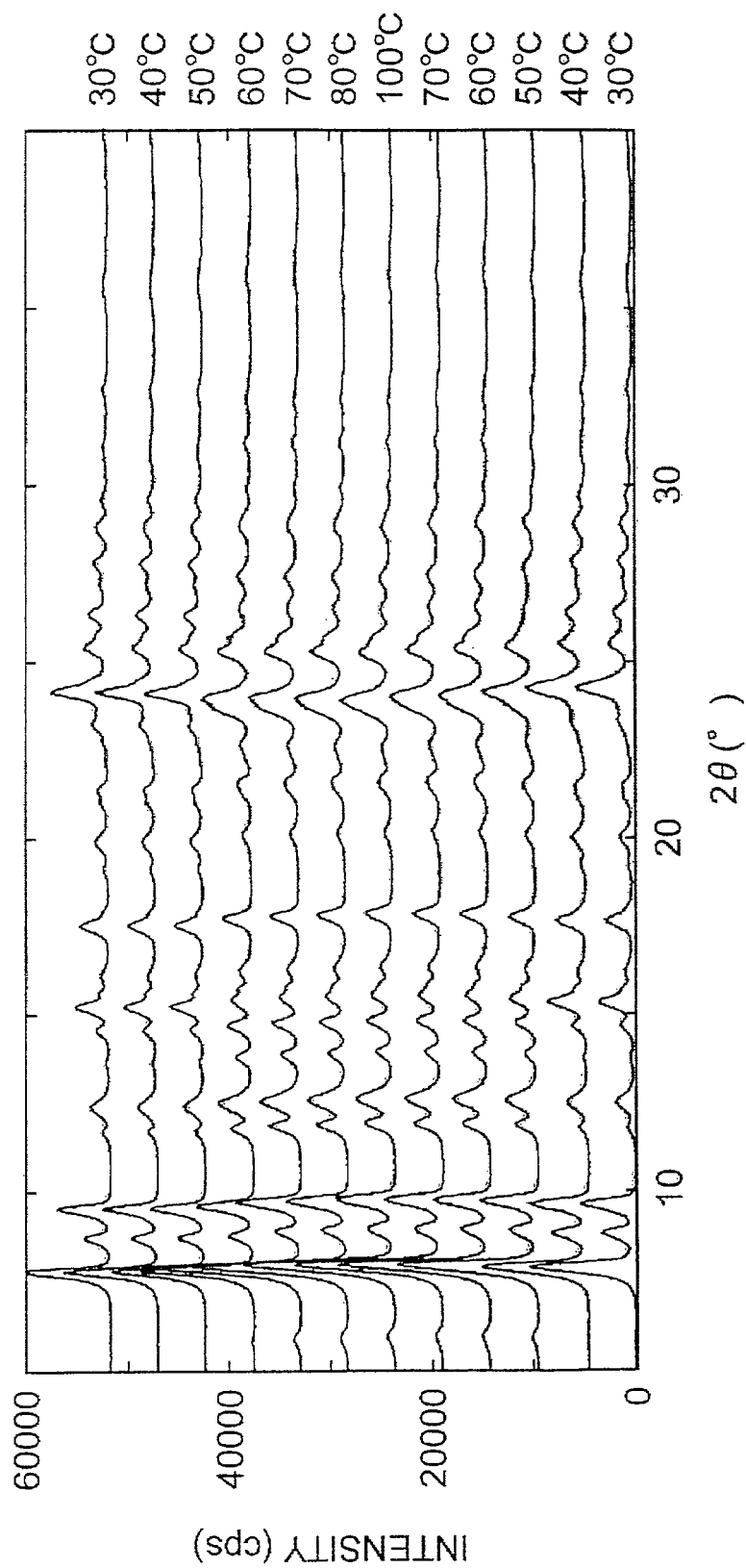
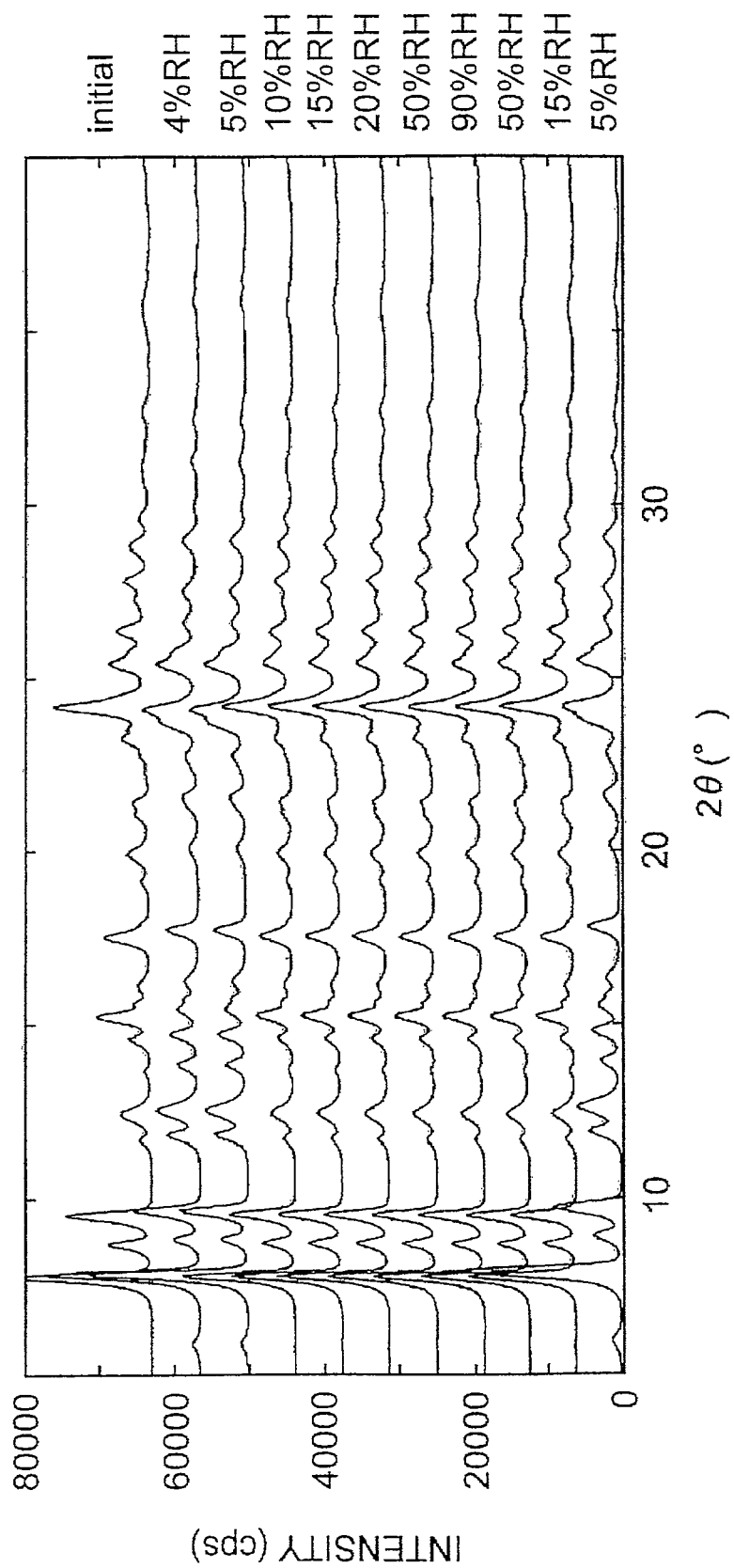




FIG. 12



# 1

## METHOD FOR PRODUCING 1,2-DIHYDROPYRIDINE-2-ONE COMPOUND

### CROSS-REFERENCES TO RELATED APPLICATIONS

This application is a divisional application of U.S. application Ser. No. 12/870,507 filed on Aug. 27, 2010, now U.S. Pat. No. 8,772,497, which is a divisional of U.S. application Ser. No. 11/649,299 filed on Jan. 4, 2007, now U.S. Pat. No. 8,304,548, which issued on Nov. 6, 2012, which is a continuation-in-part of PCT/JP2005/012364, which was filed on Jul. 5, 2005, and claims the benefit of priority of JP 2004-198709 filed in Japan on Jul. 6, 2004, and also a continuation-in-part of PCT/JP2005/012388, which was filed on Jul. 5, 2005, and claims the benefit of priority of JP 2004-198709 filed in Japan on Jul. 6, 2004. All of the contents of U.S. application Ser. No. 12/870,507, U.S. application Ser. No. 11/649,299, PCT/JP2005/012364, PCT/JP2005/012388 and JP 2004-198709 are hereby expressly incorporated by reference.

### BACKGROUND OF THE INVENTION

#### 1. Field of the Invention

The present invention relates to a method for producing a 1,2-dihydropyridine-2-one compound represented by formula (III) which comprises reacting a compound represented by formula (I) with a boronic acid derivative represented by formula (II) in the presence of a palladium compound, a copper compound, a phosphorus compound and a base. Further, the present invention relates to crystals of 3-(2-cyanophenyl)-5-(2-pyridyl)-1-phenyl-1,2-dihydropyridine-2-one and a method for producing the crystals.

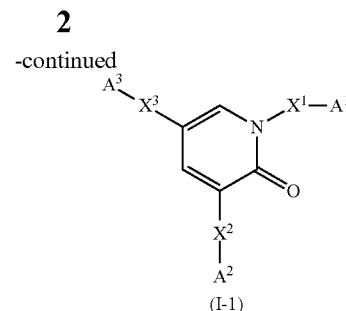
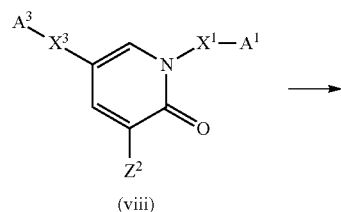
The compound of formula (III) represented by 3-(2-cyanophenyl)-5-(2-pyridyl)-1-phenyl-1,2-dihydropyridine-2-one is useful as, for example, a therapeutic agent for diseases such as Parkinson's disease, multiple sclerosis, epilepsy, etc.

#### 2. Description of the Related Art

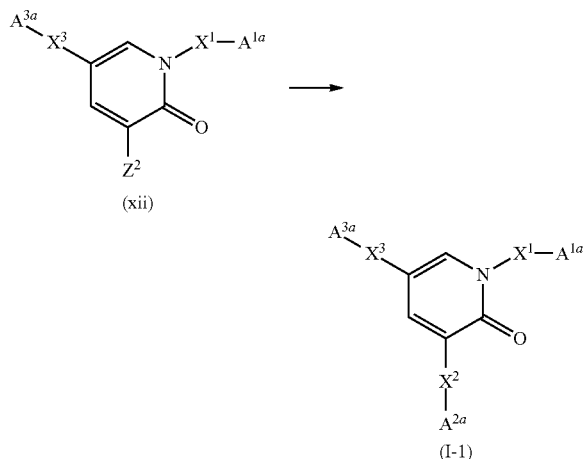
Background art concerning a method for producing the compound of formula (III) is explained below.

In the production method 2 in WO 01/96308, the coupling reaction of a compound (viii) with an arylboronic acid derivative by the use of a palladium catalyst is described as to a final step for producing a compound (I-1), but the reaction in the presence of a palladium compound, a copper compound and a phosphorus compound is neither suggested nor described which is characteristic of the present invention.

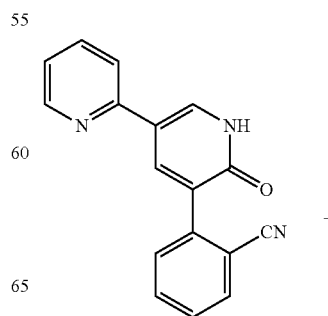
#### Production Method 2

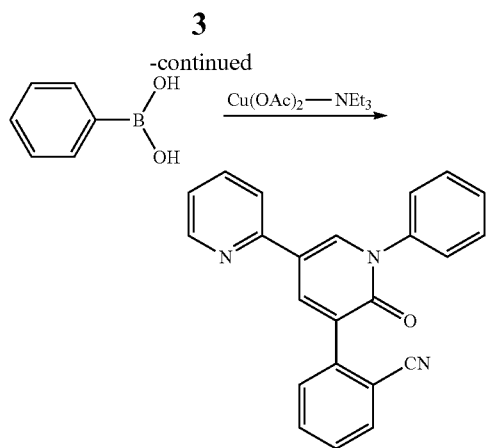


Also in the production method 3 in WO 01/96308, the coupling reaction of a compound (xii) with an arylboronic acid derivative by the use of a palladium catalyst is described as to a final step for producing a compound (I-1), but the reaction in the presence of a palladium compound, a copper compound and a phosphorus compound is neither suggested nor described which is characteristic of the present invention.



The compound of formula (III-a) described hereinafter is a well-known compound. In Example 7 in WO 01/96308, it is known that as shown in the following reaction scheme, this compound may be produced by reacting 3-(2-cyanophenyl)-5-(2-pyridyl)-2(1H)-pyridone with phenylboronic acid in the presence of copper acetate and triethylamine. But, there is neither suggested nor described a method for producing a compound of formula (III) by the reaction of a compound of formula (I) with a compound of formula (II) in the presence of a palladium compound, a copper compound, a phosphorus compound and a base which is characteristic of the present invention.





As to the compound of formula (I) represented by 3-bromo-5-(2-pyridyl)-1-phenyl-1,2-dihydropyridine-2-one (III-a), a method for producing this compound is described in claim 49 and Example 404 in WO 01/96308.

On the other hand, the effect of a copper catalyst in the Suzuki reaction is described in G. M. Boland et al., J. Chem. Soc., Perkin Trans. 1, 2591-2587 (1996). Although this reference describes "Pd(PPh<sub>3</sub>)<sub>4</sub>-CuI", copper iodide is used in a large amount of 1.1 equivalents per equivalent of a starting material in the reference. The reference neither suggests nor describes the progress of the reaction in the presence of a palladium compound, a copper compound and a phosphorus compound, in particular, the reaction in the presence of a catalytic amount of the copper compound, which is characteristic of the present invention.

On the other hand, among the compounds of formula (III), particularly, 3-(2-cyanophenyl)-5-(2-pyridyl)-1-phenyl-1,2-dihydropyridine-2-one (hereafter referred to as Compound (A) in some cases) shows significant antagonistic action against AMPA (alpha -amino-3-hydroxy-5-methyl-4-isoxazolepropionic acid) receptor (see WO 01/96308).

Although Example 7 in WO 01/96308 discloses a process for producing the compound (A), there is merely described, "the residue is purified by silica gel column chromatography (ethyl acetate/hexane=1:2)" and there is no disclosure of the form of the obtained compound.

### SUMMARY OF THE INVENTION

When in each of the methods using a palladium catalyst described as the production method 2 and production method 3 in WO 01/96308, the reaction is carried out in the presence of, for example, "palladium acetate catalyst-cesium carbonate-water", there are various problems such as the following problems: a considerable amount of compounds are produced as by-products by the cleavage of the carbon-boron bond of a compound (II) (Yoshio Urawa and three others, Pharmacia, 35(7), 706-710 (1999)) and the hydrolysis of a substituent such as a nitrile group proceeds. Therefore, an industrial method for producing a compound represented by formula (III) is desired.

Accordingly, an object of the present invention is to provide a method for industrially producing a compound of formula (III) having an excellent therapeutic effect on diseases such as Parkinson's disease, multiple sclerosis, epilepsy, etc., in good yield and high purity.

On the other hand, when a compound existing in crystal polymorphism is used as a medicament, it is necessary to stably supply the compound having uniform crystal form so that the uniform quality and the consistent potency required

for the medicament may be guaranteed. There is also a need for the crystal form capable of maintaining the same quality during its storage and its formulation process such as blending and granulation.

Since a drug substance is industrially used in a large amount, desirable crystal forms are those having high explosion concentration high limit and minimum ignition energy, index of explosiveness and dangerousness.

Generally, powders that tend to be charged have great adhesiveness to other objects; and there is concern about their adhesion to protective goods or the skin.

When a drug substance has chargeability, it happens that the production efficiency and workability lower if the compound adheres to a rotary blade at a milling stage in the manufacture of the compound, or adheres to and agglomerate on the production machines during the process of formulation. When a large quantity of powders having chargeability is processed on an industrial scale, there is the possibility that dust explosion will occur. It is, therefore, desired that a compound crystal having weak chargeability be used as the drug substance.

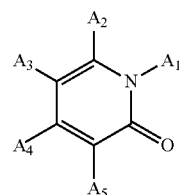
As for a compound having high pharmacological activity such as the drug substance, the standpoint of the avoidance of exposure to the workers and the prevention of the facility contamination makes powders that do not tend to be charged desirable.

For the reasons above, when the active pharmaceutical ingredient of a medicament is obtained as a crystalline substance, it desirably comprises a homogeneous crystal form, has consistently preferable properties, and does not contain impurities such as metals. There has also been a need to develop a process for stably producing such crystals on an industrial scale.

Accordingly, an object of the present invention to provide a crystal comprising a homogeneous crystal form of Compound (A) and a production process therefor.

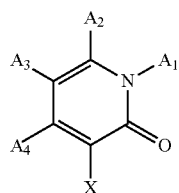
The present inventors earnestly investigated in order to solve the above problems, and consequently found the following production method of the compound of formula (III), crystal form of Compound (A) and production method of the crystal form, whereby the present invention has been accomplished. That is, the present invention relates to the following production methods 1) to 13), crystal forms 14) to 19), production methods 20) to 28), crystal forms 29) to 38), medicament 39), composition 40) and agents 41) to 50).

1) A method for producing a compound represented by formula (III):

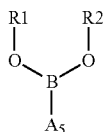


wherein A<sub>1</sub>, A<sub>2</sub>, A<sub>3</sub>, A<sub>4</sub> and A<sub>5</sub> are as defined below, or a salt thereof, which comprises reacting a compound represented by formula (I):

5



wherein each of A<sub>1</sub>, A<sub>2</sub>, A<sub>3</sub> and A<sub>4</sub>, which may be the same or different, is a hydrogen atom, an optionally substituted 6- to 14-membered aromatic hydrocarbon ring group or an optionally substituted 5- to 14-membered heteroaromatic ring group, and X is a leaving group, or a salt thereof with a compound represented by formula (II):

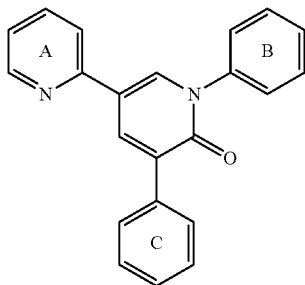


wherein A<sub>5</sub> is an optionally substituted 6- to 14-membered aromatic hydrocarbon ring group or an optionally substituted 5- to 14-membered heteroaromatic ring group; and R1 and R2 are as follows: 1) each of R1 and R2, which may be the same or different, is a hydrogen atom or a C1-6 alkyl group, and 2) the compound of formula (II) may form boroxine (a trimer) when both R1 and R2 are hydrogen atoms, or 3) R1, R2, the oxygen atoms and the boron atom, when taken together, form a 5- or 6-membered ring group optionally substituted by one to four C1-6 alkyl groups, in the presence of a palladium compound, a copper compound, a phosphorus compound and a base.

2) A production method according to 1) above, wherein each of A<sub>2</sub> and A<sub>4</sub> is a hydrogen atom.

3) A production method according to 1) or 2) above, wherein each of A<sub>1</sub>, A<sub>3</sub> and A<sub>5</sub> is a phenyl group, a pyridyl group, a pyrimidyl group, a thienyl group or a furyl group.

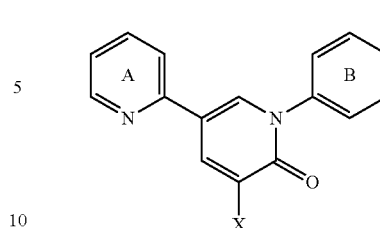
4) A production method according to any one of 1) to 3) above, wherein a compound represented by formula (III-a):



wherein the ring A, ring B and ring C are as defined below, or a salt thereof is produced by reacting a compound represented by formula (I-a):

6

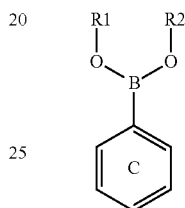
(I)



(I-a)

wherein the ring A is an optionally substituted 2-pyridyl group, the ring B is an optionally substituted phenyl group, and X is a leaving group, or a salt thereof with a compound represented by formula (II-a):

(II)

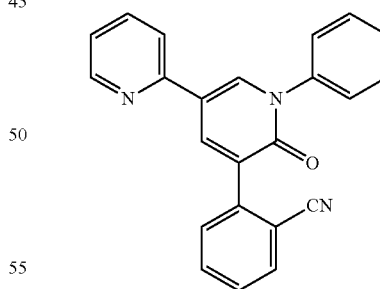


(II-a)

wherein the ring C is an optionally substituted phenyl group; and R1 and R2 are as follows: 1) each of R1 and R2, which may be the same or different, is a hydrogen atom or a C1-6 alkyl group, and 2) the compound of formula (II-a) may form boroxine (a trimer) when both R1 and R2 are hydrogen atoms, or 3) R1, R2, the oxygen atoms and the boron atom, when taken together, form a 5- or 6-membered ring group optionally substituted by one to four C1-6 alkyl groups, in the presence of a palladium compound, a copper compound, a phosphorus compound and a base.

5) A production method according to 4) above, wherein a compound represented by formula (III-b):

(III-a)

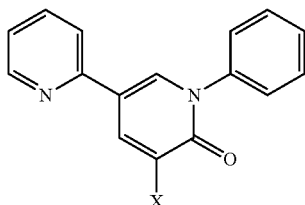


(III-b)

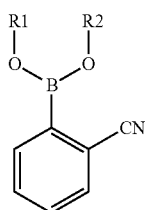
or a salt thereof is produced by reacting a compound represented by formula (I-b):

60

7

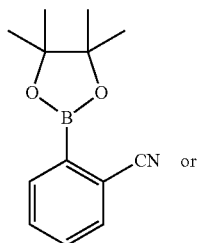
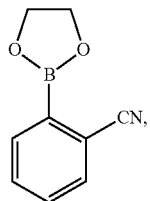
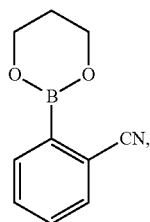


wherein X is a leaving group, or a salt thereof with a compound represented by formula (II-b):



wherein R1 and R2 are as defined above, in a solvent in the presence of a palladium compound, a copper compound, a phosphorus compound and a base.

6) A production method according to 5) above, wherein the compound (II-b) is a compound represented by formula (II-b-1), formula (II-b-2), formula (II-b-3) or formula (II-b-4):

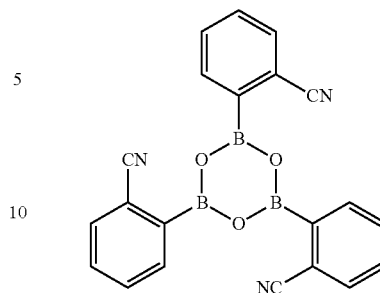


8

-continued

(II-b-4)

(I-b)



7) A production method according to any one of 1) to 6) above, wherein X is a halogen atom, an alkylsulfonyloxy group or an arylsulfonyloxy group.

8) A production method according to any one of 1) to 7) above, wherein the palladium compound is palladium acetate, palladium chloride or palladium hydroxide.

9) A production method according to any one of 1) to 8) above, wherein the phosphorus compound is triphenylphosphine or tri-tert-butylphosphine.

10) A production method according to any one of 1) to 9) above, wherein the copper compound is cuprous bromide, cuprous iodide, cuprous chloride or cuprous acetate.

11) A production method according to any one of 1) to 10) above, wherein the base is cesium carbonate, sodium carbonate or potassium carbonate.

12) A production method according to any one of 1) to 11) above, wherein the copper compound is used in an amount of 0.01 to 0.05 mole per mole of the compound represented by formula (I).

13) A production method according to any one of 1) to 12) above, wherein the reaction is carried out in a solvent and 1,2-dimethoxyethane or toluene is used as the solvent for reaction.

14) A crystal of 3-(2-cyanophenyl)-5-(2-pyridyl)-1-phenyl-1,2-dihydropyridin-2-one hydrate having a diffraction peak at a diffraction angle ( $2\theta \pm 0.2^\circ$ ) of  $8.7^\circ$  in a powder X-ray diffraction (Hydrate crystal).

15) A crystal of 3-(2-cyanophenyl)-5-(2-pyridyl)-1-phenyl-1,2-dihydropyridin-2-one hydrate having a diffraction peak at a diffraction angle ( $2\theta \pm 0.2^\circ$ ) of  $12.5^\circ$  in a powder X-ray diffraction (Hydrate crystal).

16) A crystal of 3-(2-cyanophenyl)-5-(2-pyridyl)-1-phenyl-1,2-dihydropyridin-2-one hydrate having diffraction peaks at diffraction angles ( $2\theta \pm 0.2^\circ$ ) of  $8.7^\circ$  and  $12.5^\circ$  in a powder X-ray diffraction (Hydrate crystal).

17) A crystal of 3-(2-cyanophenyl)-5-(2-pyridyl)-1-phenyl-1,2-dihydropyridin-2-one hydrate having an absorption peak at a wavenumber of  $1588 \pm 1 \text{ cm}^{-1}$  in an infrared absorption spectrum (KBr method) (Hydrate crystal).

18) A crystal of 3-(2-cyanophenyl)-5-(2-pyridyl)-1-phenyl-1,2-dihydropyridin-2-one hydrate having absorption peaks at wavenumbers of  $1588 \pm 1 \text{ cm}^{-1}$  and  $751 \pm 1 \text{ cm}^{-1}$  in an infrared absorption spectrum (KBr method) (Hydrate crystal).

18-2) The crystal according to any one of 14) to 16) above having an absorption peak at a wavenumber of  $1588 \pm 1 \text{ cm}^{-1}$  in an infrared absorption spectrum (KBr method) (Hydrate crystal).

18-3) The crystal according to any one of 14) to 16) above having absorption peaks at wavenumbers of  $1588 \pm 1 \text{ cm}^{-1}$  and  $751 \pm 1 \text{ cm}^{-1}$  in an infrared absorption spectrum (KBr method) (Hydrate crystal).

18-4) The crystal according to any one of 14) to 18), 18-2) and 18-3) above having a palladium content of 20 ppm or less, preferably 15 ppm or less (Hydrate crystal).

19) A crystal of 3-(2-cyanophenyl)-5-(2-pyridyl)-1-phenyl-1,2-dihydropyridin-2-one hydrate having peaks at chemical shifts of around 146.7 ppm and around 123.3 ppm in a  $^{13}\text{C}$  Solid State Nuclear Magnetic Resonance spectrum (Hydrate crystal).

20) A process for producing a crystal of 3-(2-cyanophenyl)-5-(2-pyridyl)-1-phenyl-1,2-dihydropyridin-2-one hydrate according to any one of 14) to 18), 18-1), 18-2), 18-3), 18-4) and 19) above, the process comprising the step of crystallizing 3-(2-cyanophenyl)-5-(2-pyridyl)-1-phenyl-1,2-dihydropyridin-2-one with an aid of one or two crystallization solvents selected from the group consisting of an alcoholic solvent, an alkylketone solvent, and water.

21) The process according to 20) above, wherein the crystallization solvent is a mixed solvent of acetone and water.

22) The process according to 20) above, wherein the crystallization solvent is a mixed solvent of acetone and water with a volume ratio of 37:3 to 24:16, preferably a mixed solvent of acetone and water with a volume ratio of 9:1 to 7:3, and more preferably a mixed solvent of acetone and water with a volume ratio of 8:2 formed by dissolving the crystals in a mixed solvent of acetone and water with a volume ratio of 9:1 and thereafter adding water to the mixed solvent.

23) The process according to any one of 20) to 22) above, wherein the crystallization is carried out at a temperature of 60 to  $-30^\circ\text{C}$ .

24) The process according to any one of 20) to 22) above comprising the steps of heating a solution of 3-(2-cyanophenyl)-5-(2-pyridyl)-1-phenyl-1,2-dihydropyridin-2-one dissolved in the crystallization solvent at a temperature of  $50^\circ\text{C}$ . or more (preferably at a temperature of the reflux temperature of the crystallization solvent to  $50^\circ\text{C}$ ., more preferably at a temperature of  $65$  to  $55^\circ\text{C}$ .) and thereafter cooling the solution to a temperature of 10 to  $-20^\circ\text{C}$ . (preferably to a temperature of 10 to  $5^\circ\text{C}$ .) at a cooling rate of 40 to  $5^\circ\text{C}$ . per hour (preferably at a cooling rate of 25 to  $15^\circ\text{C}$ . per hour).

25) The process according to any one of 20) to 24) above, wherein the crystallization solvent is used in a volume ratio of 10- to 50-fold (v/w) based on the weight of 3-(2-cyanophenyl)-5-(2-pyridyl)-1-phenyl-1,2-dihydropyridin-2-one. The quantity of the crystallization solvent is preferably from 30- to 50-fold (v/w), more preferably about 40-fold (v/w) where acetone and water (9:1) is used as the crystallization solvent and about 45-fold (v/w) where acetone and water (8:2) is used as the crystallization solvent.

26) The process according to any one of 20) to 25) above, wherein seed crystals (a small amount of crystals of 3-(2-cyanophenyl)-5-(2-pyridyl)-1-phenyl-1,2-dihydropyridin-2-one hydrate) are added at a temperature of  $60^\circ\text{C}$ . or less (preferably at a temperature of 55 to  $0^\circ\text{C}$ ., more preferably 55 to  $35^\circ\text{C}$ ., and most preferably about  $40^\circ\text{C}$ .).

27) The process according to any one of 20) to 26) above, wherein the crystals are dried under reduced pressure after the crystallization.

28) The process according to any one of 20) to 27) above, wherein the crystals are allowed to stand in the atmosphere after the crystallization and the drying under reduced pressure.

28-1) The process according to any one of 20) to 26) above, wherein the crystals are allowed to stand in the atmosphere after the crystallization.

28-2) The process according to 27) above, wherein the crystals are allowed to stand in the atmosphere after the drying under reduced pressure.

29) An anhydrous crystal of 3-(2-cyanophenyl)-5-(2-pyridyl)-1-phenyl-1,2-dihydropyridin-2-one having a diffraction peak at a diffraction angle ( $2\theta \pm 0.2^\circ$ ) of  $10.3^\circ$  in a powder X-ray diffraction (Anhydrous crystal form I).

30) The crystal according to 29) above further having a diffraction peak at a diffraction angle ( $2\theta \pm 0.2^\circ$ ) of  $19.1^\circ$  in a powder X-ray diffraction (Anhydrous crystal form I).

31) An anhydrous crystal of 3-(2-cyanophenyl)-5-(2-pyridyl)-1-phenyl-1,2-dihydropyridin-2-one having peaks at chemical shifts of around 149.0 ppm and around 125.6 ppm in a  $^{13}\text{C}$  Solid State Nuclear Magnetic Resonance spectrum (Anhydrous crystal form I).

32) An anhydrous crystal of 3-(2-cyanophenyl)-5-(2-pyridyl)-1-phenyl-1,2-dihydropyridin-2-one having a diffraction peak at a diffraction angle ( $2\theta \pm 0.2^\circ$ ) of  $16.7^\circ$  in a powder X-ray diffraction (Anhydrous crystal form V).

33) The crystal according to 32) above further having diffraction peaks at diffraction angles ( $2\theta \pm 0.2^\circ$ ) of  $12.9^\circ$  and  $24.9^\circ$  in a powder X-ray diffraction (Anhydrous crystal form V).

34) An anhydrous crystal of 3-(2-cyanophenyl)-5-(2-pyridyl)-1-phenyl-1,2-dihydropyridin-2-one having an absorption peak at a wavenumber of  $1658 \pm 1\text{ cm}^{-1}$  in an infrared absorption spectrum (KBr method) (Anhydrous crystal form V).

35) The crystal according to 34) above further having an absorption peak at a wavenumber of  $501 \pm 1\text{ cm}^{-1}$  in an infrared absorption spectrum (KBr method) (Anhydrous crystal form V).

36) An anhydrous crystal of 3-(2-cyanophenyl)-5-(2-pyridyl)-1-phenyl-1,2-dihydropyridin-2-one having peaks at chemical shifts of around 145.9 ppm and around 137.7 ppm in a  $^{13}\text{C}$  Solid State Nuclear Magnetic Resonance spectrum (Anhydrous crystal form V).

37) An anhydrous crystal of 3-(2-cyanophenyl)-5-(2-pyridyl)-1-phenyl-1,2-dihydropyridin-2-one having diffraction peaks at diffraction angles ( $2\theta \pm 0.2^\circ$ ) of  $23.7^\circ$  and  $25.0^\circ$  in a powder X-ray diffraction (Anhydrous crystal form III).

38) The crystal according to 32) above further having diffraction peaks at diffraction angles ( $2\theta \pm 0.2^\circ$ ) of  $5.7^\circ$  and  $9.5^\circ$  in a powder X-ray diffraction (Anhydrous crystal form III).

39) A medicament comprising the crystal according to 14) above.

40) A pharmaceutical composition comprising the crystal according to 14) above.

41) A therapeutic or prophylactic agent for an acute neurodegenerative disease comprising the crystal according to 14) above.

42) A therapeutic or prophylactic agent for neuropathy caused by acute phase of cerebrovascular disorder, head injury, spinal cord injury or hypoxia, or neuropathy caused by hypoglycemia, comprising the crystal according to 14) above.

43) A therapeutic or prophylactic agent for a chronic neurodegenerative disease comprising the crystal according to 14) above.

44) A therapeutic or prophylactic agent for Alzheimer's disease, Parkinson's disease, Huntington's chorea, amyotrophic lateral sclerosis or spinocerebellar degeneration, comprising the crystal according to 14) above.

45) A therapeutic or prophylactic agent for epilepsy, hepatic encephalopathy, peripheral neuropathy, Parkinsonism, spastic paralysis, pain, neuralgia, schizophrenia, anxiety, drug-dependence, nausea, vomiting, dysuria, vision impairment caused by glaucoma, hearing impairment caused by antibiotics, or food poisoning, the agent comprising the crystal according to 14) above.

46) A therapeutic or prophylactic agent for infectious encephalomyelitis, cerebrovascular dementia, or dementia or neurological symptom caused by meningitis, comprising the crystal according to 14) above.

47) A therapeutic or prophylactic agent for a demyelinating disease comprising the crystal according to 14) above.

48) The therapeutic or prophylactic agent according to 47) above, wherein the infectious encephalomyelitis is HIV encephalomyelitis.

49) The therapeutic or prophylactic agent according to 47) above, wherein the demyelinating disease is encephalitis, acute sporadic encephalomyelitis, multiple sclerosis, acute polyradiculoneuropathy, Guillain-Barre syndrome, chronic inflammatory demyelinating polyradiculoneuropathy, Marchifava-Bignami disease, central pontomedullary myelinolysis, neuromyelitis optica, Devic's disease, Balo's disease, HIV-associated myelopathy, HTLV-associated myelopathy, progressive multifocal leukoencephalitis or a secondary demyelinating disease.

50) The therapeutic or prophylactic agent according to 49) above, wherein the secondary demyelinating disease is CNS lupus erythematosus, polyarteritis nodosa, Sjogren's syndrome, sarcoidosis or dissociated cerebral vasculitis.

#### BRIEF DESCRIPTION OF THE DRAWINGS

FIG. 1 shows an infrared spectrum (KBr method) of the crystals obtained in Example 5.

FIG. 2 shows an infrared spectrum (KBr method) of the crystals obtained in Example 6.

FIG. 3 shows a powder X-ray diffraction pattern of the crystals obtained in Reference Example 1.

FIG. 4 shows a powder X-ray diffraction pattern of the crystals obtained in Example 5.

FIG. 5 shows a powder X-ray diffraction pattern of the crystals obtained in Example 6.

FIG. 6 shows a powder X-ray diffraction pattern of the crystals obtained in Example 7.

FIG. 7 shows a powder X-ray diffraction pattern of the crystals as described in Example E1.

FIG. 8 shows a  $^{13}\text{C}$  Solid State Nuclear Magnetic Resonance (NMR) spectrum of the crystals obtained in Example 5.

FIG. 9 shows a  $^{13}\text{C}$  Solid State NMR spectrum of the crystals obtained in Example 7.

FIG. 10 shows a  $^{13}\text{C}$  Solid State NMR spectrum of the crystals obtained in Example 6.

FIG. 11 shows powder X-ray diffraction patterns of Hydrate crystal at various temperatures.

FIG. 12 shows powder X-ray diffraction patterns of Hydrate crystal under various relative humidities.

#### DESCRIPTION OF THE PREFERRED EMBODIMENTS

The symbols and terms used in the present specification are explained below.

The term "6- to 14-membered aromatic hydrocarbon ring group" means an aromatic hydrocarbon ring group comprising 6 to 14 carbon atoms and also includes fused-ring groups such as monocyclic groups, bicyclic groups, tricyclic groups, etc. Specific examples of said group are phenyl group, indenyl group, 1-naphthyl group, 2-naphthyl group, azulenyl group, heptalenyl group, biphenyl group, indacenyl group, acenaphthyl group, fluorenyl group, phenalenyl group, phenanthrenyl group, anthracenyl group, etc.

The term "5- to 14-membered heteroaromatic ring group" means a monocyclic, bicyclic or tricyclic 5- to 14-membered

heteroaromatic ring group containing one or more heteroatoms selected from the group consisting of nitrogen atom, sulfur atom and oxygen atom. Specific examples of said group are 1) nitrogen-containing heteroaromatic ring groups such as pyrrolyl group, pyridyl group, pyridazinyl group, pyrimidinyl group, pyrazinyl group, triazolyl group, tetrazolyl group, benzotriazolyl group, pyrazolyl group, imidazolyl group, benzimidazolyl group, indolyl group, isoindolyl group, indolizynyl group, purinyl group, indazolyl group, quinolyl group, isoquinolyl group, quinolizynyl group, phthalazyl group, naphthyridinyl group, quinoxalyl group, quinazolinyl group, cinnolynyl group, pteridinyl group, imidazotriazinyl group, pyrazinopyridazinyl group, acridinyl group, phenanthridinyl group, carbazolyl group, carbazolinyl group, perimidinyl group, phenanthrolinyl group, phenazinyl group, imidazopyridinyl group, imidazopyrimidinyl group, pyrazolopyridinyl group, etc., 2) sulfur-containing heteroaromatic ring groups such as thienyl group, benzothienyl group, etc., 3) oxygen-containing heteroaromatic ring groups such as furyl group, pyranyl group, cyclopentapyranyl group, benzofuryl group, isobenzofuryl group, etc., and 4) heteroaromatic ring groups containing two or more heteroatoms of different kinds, such as thiazolyl group, isothiazolyl group, benzothiazolyl group, benzthiadiazolyl group, phenothiazinyl group, isoxazolyl group, furazanyl group, phenoxazinyl group, oxazolyl group, isoxazolyl group, benzoxazolyl group, oxadiazolyl group, pyrazoloxazolyl group, imidazothiazolyl group, thienofuranyl group, furopyrrolyl group, pyridoxazinyl group, etc.

Each of  $A_1$ ,  $A_2$ ,  $A_3$  and  $A_4$  is a hydrogen atom, an optionally substituted 6- to 14-membered aromatic hydrocarbon ring group or an optionally substituted 5- to 14-membered heteroaromatic ring group. More preferably, each of  $A_2$  and  $A_4$  is a hydrogen atom and each of  $A_1$  and  $A_3$  is an optionally substituted 6- to 14-membered aromatic hydrocarbon ring group or an optionally substituted 5- to 14-membered heteroaromatic ring group. Most preferably, each of  $A_1$  and  $A_3$  is, for example, an optionally substituted phenyl, pyridyl, pyrimidinyl, thienyl or furyl group.

$A_5$  is an optionally substituted 6- to 14-membered aromatic hydrocarbon ring group or an optionally substituted 5- to 14-membered heteroaromatic ring group.  $A_5$  is more preferably, for example, an optionally substituted phenyl, pyrrolyl, pyridyl, pyridazinyl, pyrimidinyl, pyrazinyl, thienyl, thiazolyl, furyl, naphthyl, quinolyl, isoquinolyl, indolyl, benzimidazolyl, benzothiazolyl, benzoxazolyl, imidazopyridyl or pyrrolidinyl group.  $A_5$  is most preferably, for example, an optionally substituted phenyl, pyridyl, pyrimidinyl, thienyl or furyl group.

When the group represented by any of  $A_1$ ,  $A_2$ ,  $A_3$ ,  $A_4$  and  $A_5$  in the above formula is an optionally substituted 6- to 14-membered aromatic hydrocarbon ring group or an optionally substituted 5- to 14-membered heteroaromatic ring group, it may have one to four substituents which may be the same or different and are selected from the following substituents.

In the above formula, the ring A is an optionally substituted 2-pyridyl group and each of the ring B and the ring C is an optionally substituted phenyl group. The ring A, ring B and ring C may also have one to four substituents which may be the same or different and are selected from the following substituents.

The substituents include, for example, hydroxyl group, nitrile groups, halogen atoms, C1-6 alkyl groups, C2-6 alkenyl groups, C2-6 alkynyl groups, C3-8 cycloalkyl groups, C1-6 alkoxy groups, C1-6 alkylthio groups, C1-6 alkoxy carbonyl groups, C1-6 alkanoyl groups (C1-6 alkylcarbonyl

groups), C1-6 alkylsulfonyl groups, amino group optionally substituted by a C1-6 alkyl group, amino group optionally substituted by a formyl group, amino group optionally substituted by a C1-6 alkanoyl group, amino group optionally substituted by a C1-6 alkylsulfonyl group, carbamoyl group optionally substituted by one or two C1-6 alkyl groups, and C1-6 alkoxyimino groups. Of these, the nitrile groups and halogen atoms are preferable.

The term "halogen atoms" means a fluorine atom, chlorine atom, bromine atom, iodine atom and the like. The halogen atoms are preferably a chlorine atom and a bromine atom.

The term "C1-6 alkyl groups" means alkyl groups of 1 to 6 carbon atoms. Preferable examples of these groups are linear or branched alkyl groups such as methyl group, ethyl group, n-propyl group, i-propyl group, n-butyl group, i-butyl group, tert-butyl group, n-pentyl group, i-pentyl group, neopentyl group, n-hexyl group, 1-methylpropyl group, 1,2-dimethylpropyl group, 2-ethylpropyl group, 1-methyl-2-ethylpropyl group, 1-ethyl-2-methylpropyl group, 1,1,2-trimethylpropyl group, 1-methylbutyl group, 2-methylbutyl group, 1,1-dimethylbutyl group, 2,2-dimethylbutyl group, 2-ethylbutyl group, 1,3-dimethylbutyl group, 2-methylpentyl group, 3-methylpentyl group, etc.

The term "C2-6 alkenyl groups" means alkenyl groups of 2 to 6 carbon atoms. Preferable examples of these groups are linear or branched alkenyl groups such as vinyl group, allyl group, 1-propenyl group, isopropenyl group, 1-buten-1-yl group, 1-buten-2-yl group, 1-buten-3-yl group, 2-buten-1-yl group, 2-buten-2-yl group, etc.

The term "C2-6 alkynyl groups" means alkynyl groups of 2 to 6 carbon atoms. Preferable examples of these groups are linear or branched alkynyl groups such as ethynyl group, 1-propynyl group, 2-propynyl group, butynyl group, pentynyl group, hexynyl group, etc.

The term "C3-8 cycloalkyl groups" means cyclic alkyl groups of 3 to 8 carbon atoms. Preferable examples of these groups are cyclopropyl group, cyclobutyl group, cyclopentyl group, cyclohexyl group, cycloheptyl group, cyclooctyl group, etc.

The term "C1-6 alkoxy groups" means groups formed by the replacement of a hydrogen atom of an alkyl group of 1 to 6 carbon atoms by an oxygen atom. Preferable examples of said groups are methoxy group, ethoxy group, n-propoxy group, i-propoxy group, sec-propoxy group, n-butoxy group, i-butoxy group, sec-butoxy group, tert-butoxy group, n-pentyloxy group, i-pentyloxy group, sec-pentyloxy group, tert-pentyloxy group, n-hexyloxy group, i-hexyloxy group, 1,2-dimethylpropoxy group, 2-ethylpropoxy group, 1-methyl-2-ethylpropoxy group, 1-ethyl-2-methylpropoxy group, 1,1,2-trimethylpropoxy group, 1,1-dimethylbutoxy group, 2,2-dimethylbutoxy group, 2-ethylbutoxy group, 1,3-dimethylbutoxy group, 2-methylpentyloxy group, 3-methylpentyloxy group, hexyloxy group, etc.

The term "C1-6 alkylthio groups" means groups formed by the replacement of a hydrogen atom of an alkyl group of 1 to 6 carbon atoms by a sulfur atom. Preferable examples of said groups are methylthio group, ethylthio group, n-propylthio group, i-propylthio group, n-butylthio group, i-butylthio group, tert-butylthio group, n-pentylthio group, i-pentylthio group, neopentylthio group, n-hexylthio group, 1-methylpropylthio group, etc.

The term "C1-6 alkoxy carbonyl groups" means groups formed by bonding of a carbonyl group to any of the above-exemplified alkoxy groups. Preferable examples of said groups are methoxycarbonyl group, ethoxycarbonyl group, etc.

The term "C1-6 alkanoyl groups (C1-6 alkylcarbonyl groups)" means groups formed by the replacement of a hydrogen atom of an alkyl group of 1 to 6 carbon atoms by a carbonyl group. Preferable examples of said groups are acetyl group, propionyl group, butyryl group, etc.

The term "C1-6 alkylsulfonyl groups" means groups formed by the replacement of a hydrogen atom of an alkyl group of 1 to 6 carbon atoms by a sulfonyl group. Preferable examples of said groups are methanesulfonyl group, ethanesulfonyl group, etc.

The term "amino group optionally substituted by a C1-6 alkyl group" means an amino group that may have an alkyl group of 1 to 6 carbon atoms bonded thereto. Preferable examples of such an amino group are amino group, methylamino group, ethylamino group, propylamino group, etc.

"Amino group optionally substituted by a formyl group" includes, for example, amino group, formylamino group, etc.

The term "amino group optionally substituted by a C1-6 alkanoyl group" means an amino group that may have an alkanoyl group of 1 to 6 carbon atoms bonded thereto. Preferable examples of such an amino group are acetilamino group, propionylamino group, butyrylamino group, etc.

The term "amino group optionally substituted by a C1-6 alkylsulfonyl group" means an amino group that may have an alkylsulfonyl group of 1 to 6 carbon atoms bonded thereto. Preferable examples of such an amino group are amino group, methanesulfonylamino group, ethanesulfonylamino group, n-propanesulfonylamino group, n-butanesulfonylamino group, N-methylmethanesulfonylamino group, etc.

The term "carbamoyl group optionally substituted by one or two C1-6 alkyl groups" means a carbamoyl group one or two hydrogen atoms of which may be replaced by one or two, respectively, C1-6 alkyl groups. Preferable examples of said groups are N-methylcarbamoyl group, N,N-dimethylcarbamoyl group, N-ethylcarbamoyl group, N,N-diethylcarbamoyl group, etc.

The term "C1-6 alkoxyimino groups" means groups formed by the replacement of a hydrogen atom of an imino group by a C1-6 alkoxy group. Preferable examples of said groups are methoxyimino group, ethoxyimino group, etc.

The passage "X is a leaving group" means that X is a halogen atom, an alkylsulfonyloxy group or an arylsulfonyloxy group.

The passage "X is a halogen atom, an alkylsulfonyloxy group or an arylsulfonyloxy group" means that X is a halogen atom such as fluorine atom, chlorine atom, bromine atom or iodine atom; an alkylsulfonyloxy group such as trifluoromethanesulfonyloxy group; or an arylsulfonyloxy group such as phenylsulfonyloxy group. X is preferably a halogen atom such as chlorine atom or bromine atom, or an alkylsulfonyloxy group such as trifluoromethanesulfonyloxy group.

The sentence "R1 and R2 are as follows: 1) each of R1 and R2, which may be the same or different, is a hydrogen atom or a C1-6 alkyl group, and 2) the compound (II) may form boroxine (a trimer) when both R1 and R2 are hydrogen atoms, or 3) R1, R2, the oxygen atoms and the boron atom, when taken together, form a 5- or 6-membered ring group optionally substituted by one to four C1-6 alkyl groups" in the case of the compound (II) means that the compound (II) is, for example, a phenylboronic acid derivative in which the hydrogen atom of the hydroxyl group may be replaced by a C1-6 alkyl group; a 2-phenyl-[1,3,2]-dioxaboronate the ring-forming methylene groups of which may be substituted by one to four C1-6 alkyl groups; or a 2-phenyl-[1,3,2]-dioxaboronate derivative the ring-forming methylene groups of which may be substituted by one to four C1-6 alkyl groups.



In particular, the passage "the compound (II) may form boroxine (a trimer) when both R1 and R2 are hydrogen atoms" means that when both R1 and R2 are hydrogen atoms, the compound (II) may be a monomer or may form a cluster such as a dimer or boroxine (a trimer).

The term "a palladium compound, a copper compound and a phosphorus compound" means a combination of a palladium compound selected from the palladium compounds described hereinafter, a copper compound selected from the copper compounds described hereinafter, and a phosphorus compound selected from the phosphorus compounds described hereinafter.

The compound (I-a) is included in the compound represented by formula (I) and corresponds to a compound of formula (I) in which A<sub>1</sub> is an optionally substituted phenyl group, each of A<sub>2</sub> and A<sub>4</sub> is a hydrogen atom, and A<sub>3</sub> is an optionally substituted 2-pyridyl group.

The compound (I-b) is included in the compound represented by formula (I-a) and corresponds to a compound of formula (I-a) in which A<sub>1</sub> is a phenyl group, each of A<sub>2</sub> and A<sub>4</sub> is a hydrogen atom, and A<sub>3</sub> is a 2-pyridyl group.

The compound (II-a) is included in the compound represented by formula (II) and corresponds to a compound of formula (II) in which A<sub>5</sub> is an optionally substituted phenyl group.

The compound (II-b) is included in the compound represented by formula (II-a) and corresponds to a compound of formula (II-a) in which A<sub>5</sub> is a 2-cyanophenyl group.

The compounds (II-b-1), (II-b-2), (II-b-3) and (II-b-4) are included in the compound represented by formula (II-b). Each of the compounds (II-b-1), (II-b-2) and (II-b-3) corresponds to a compound of formula (II-b) in which R1, R2, the oxygen atoms and the boron atom are taken together to form a 5- or 6-membered ring group optionally substituted by one to four C1-6 alkyl groups. The compound (II-b-4) corresponds to boroxine (a trimer) formed by a compound of formula (II-b) in which both R1 and R2 are hydrogen atoms.

The compound (III-a) is included in the compound represented by formula (III) and corresponds to a compound of formula (III) in which each of A<sub>1</sub> and A<sub>5</sub> is an optionally substituted phenyl group, each of A<sub>2</sub> and A<sub>4</sub> is a hydrogen atom, and A<sub>3</sub> is an optionally substituted 2-pyridyl group.

The compound (III-b) is included in the compound represented by formula (III-a) and corresponds to a compound of formula (III-a) in which the ring A is a 2-pyridyl group, the ring B is a phenyl group and the ring C is a 2-cyanophenyl group.

The term "A crystal of 3-(2-cyanophenyl)-5-(2-pyridyl)-1-phenyl-1,2-dihydropyridin-2-one hydrate" is a crystal form of 3-(2-cyanophenyl)-5-(2-pyridyl)-1-phenyl-1,2-dihydropyridin-2-one containing water in the crystal, and as such, the amount of water contained in the crystal form is not particularly limited; it may be devoid of a portion of the water in the crystal. The term also encompasses the form wherein the water may coexist with adhesion water.

The term "crystal of 3-(2-cyanophenyl)-5-(2-pyridyl)-1-phenyl-1,2-dihydropyridin-2-one hydrate" means such a crystal form that it has preferably 1/2 to one molecule of water per one molecule of 3-(2-cyanophenyl)-5-(2-pyridyl)-1-phenyl-1,2-dihydropyridin-2-one in the crystal, may further contain 0 to 1/4 molecules of adhesion water and may even be devoid of 0 to 1/2 molecules of water in the crystal.

Specifically, it means the following:

(1) A crystal of 3-(2-cyanophenyl)-5-(2-pyridyl)-1-phenyl-1,2-dihydropyridin-2-one 3/4 hydrate;

(2) A crystal of 3-(2-cyanophenyl)-5-(2-pyridyl)-1-phenyl-1,2-dihydropyridin-2-one monohydrate (devoid of 1/4 in the crystal);

(3) A crystal of 3-(2-cyanophenyl)-5-(2-pyridyl)-1-phenyl-1,2-dihydropyridin-2-one 1/2 hydrate coexisting with 1/4 adhesion water; and

(4) A crystal of 3-(2-cyanophenyl)-5-(2-pyridyl)-1-phenyl-1,2-dihydropyridin-2-one 1/2 hydrate; and

(5) A crystal of 3-(2-cyanophenyl)-5-(2-pyridyl)-1-phenyl-1,2-dihydropyridin-2-one monohydrate.

The production method of the present invention is explained below in detail.

A Method for Producing a Compound of Formula (III) Represented by 3-(2-cyanophenyl)-5-(2-pyridyl)-1-phenyl-1,2-dihydropyridine-2-one (III-a)

This production method is characterized by converting a compound of formula (I) to the compound of formula (III) by reacting the compound of formula (I) with a compound of formula (II) in a solvent in the presence of a palladium compound, a copper compound and a phosphorus compound.

This reaction may be carried out also in a stream or atmosphere of an inert gas such as nitrogen, argon or the like.

As the compound (I), there can be used compounds producible by the method described in the production example 2 described hereinafter and "Chemical Society of Japan, Jikken Kagaku Koza (Experimental Chemistry) 19, 4th ed., Organic Synthesis I-Carbon Compounds. Halogen Compounds-", Maruzen Co., Ltd., Jun. 5, 1992, p 363-482, well-known compounds, purchasable compounds, and compounds easily producible from a purchasable compound by a method conventionally adopted by those skilled in the art.

As the compound (II), there can be used compounds producible by the method described in F. R. Bean et al., J. Am. Chem. Soc., 54, 4415 (1932), J. M. Sugihara et al., J. Am. Chem. Soc., 80, 2443 (1958), or the like, well-known compounds, purchasable compounds, and compounds easily producible from a purchasable compound by a method conventionally adopted by those skilled in the art.

The reaction is preferably carried out in a solvent. The solvent for reaction used is not particularly limited so long as it dissolves the starting materials to a certain degree and does not inhibit the reaction. As the solvent, there can be used, for example, organic solvents including ether solvents (e.g. tetrahydrofuran, 1,2-dimethoxyethane, diethyl ether and dioxane), aromatic hydrocarbon solvents (e.g. benzene, toluene and xylene), amide solvents (e.g. N,N-dimethylformamide, N,N-dimethylacetamide and N-methylpyrrolidone), dimethyl sulfoxide, etc.; and mixtures of any of these organic solvents and water. The solvent is suitably, for example, 1,2-dimethoxyethane or toluene.

The above term "palladium compound" means, for example, tetrakis(triphenylphosphine)palladium, tris(dibenzylideneacetone)dipalladium, bis(dibenzylideneacetone)palladium, tetrakis(tri-tert-butylphosphine)palladium, palladium acetate, dichlorobis(triphenylphosphine)palladium, dichlorobis(tri-o-tolylphosphine)palladium, dichlorobis(tricyclohexylphosphine)palladium, 1,1'-bis(diphenylphosphino)ferrocenedichloropalladium, palladium chloride, palladium hydroxide, palladium nitrate, di-μ-chlorobis(η-allyl)palladium, bis(acetyl-acetonato)palladium, dichlorobis(benzonitrile)palladium, dichlorobis(acetonitrile)palladium or the like. The palladium compound is suitably palladium acetate, palladium chloride, palladium hydroxide or the like.

The above term "copper compound" means cuprous fluoride, cuprous chloride, cuprous bromide, cuprous iodide,

cuprous acetate or the like. The copper compound is suitably cuprous bromide, cuprous iodide, cuprous chloride or cuprous acetate.

The above term "phosphorus compound" means, for example, triphenylphosphine, tri(2-methylphenyl)phosphine, bis(diphenylphosphino)methane, bis(diphenylphosphino)ethane, bis(diphenylphosphino)propane, bis(diphenylphosphino)butane, bis(diphenylphosphino)pentane, bis(diphenylphosphino)hexane, 2,2'-bis(diphenylphosphino)-1,1'-binaphthyl, tri-tert-butylphosphine, tri(4-methylphenyl)phosphine, tricyclohexylphosphine, 2-(di-tert-butylphosphino)biphenyl, 2-(dicyclohexylphosphino)biphenyl, 1,1'-bis(diphenylphosphino)ferrocene or the like. The phosphorus compound is suitably, for example, triphenylphosphine, tri-tert-butylphosphine or tri(4-methylphenyl)phosphine, more suitably triphenylphosphine or tri-tert-butylphosphine.

The above term "base" means an inorganic base such as sodium hydroxide, barium hydroxide, sodium carbonate, potassium carbonate, cesium carbonate, sodium hydrogen carbonate, potassium hydrogencarbonate, potassium phosphate, cesium fluoride, potassium fluoride or the like; an alkali metal alkoxide such as sodium ethoxide, sodium tert-butoxide, potassium tert-butoxide or the like; or an organic amine such as N-methylmorpholine, N,N-dimethylaniline, DBU, triethylamine or the like. The base is suitably, for example, sodium carbonate, potassium carbonate, cesium carbonate, sodium hydrogencarbonate or potassium hydrogencarbonate, more suitably sodium carbonate, potassium carbonate or cesium carbonate.

The reaction temperature is usually varied depending on the starting materials, the solvent and other reagents used in the reaction and is suitably 100° C. to 50° C. (the internal temperature of a reactor), more suitably 90° C. to 60° C. (the internal temperature of the reactor).

The reaction time is usually varied depending on the starting materials, the solvent, other reagents used in the reaction and the reaction temperature. It is suitable to conduct stirring for 1 to 10 hours, more suitably about 4 hours, in the above reaction temperature range after the addition of the reagents.

The compound (II) may be used in an amount of 1 to 10 moles, suitably 1 to 3 moles, more suitably 1.5 moles, per mole of the compound (I).

The above-mentioned palladium compound may be used in an amount of 0.001 to 0.1 mole, suitably 0.01 to 0.05 mole, more suitably 0.02 mole, per mole of the compound (I).

The above-mentioned copper compound may be used in an amount of 0.001 to 0.2 mole, suitably 0.01 to 0.1 mole, more suitably 0.05 mole, per mole of the compound (I).

The above-mentioned phosphorus compound may be used in an amount of 0.001 to 0.4 mole, suitably 0.01 to 0.2 mole, more suitably 0.05 to 0.1 mole, per mole of the compound (I).

The above-mentioned base may be used in an amount of 1 to 10 moles, suitably 1 to 5 moles, more suitably 1.5 moles, per mole of the compound (I).

It is known that in the Suzuki coupling, the addition of water to a reaction system gives a good result ("Efficient Synthesis of Losartan, A Nonpeptide Angiotensin II Receptor Antagonist", Robert D. Larsen et al., J. Org. Chem., 1994, 59, 6391-6394, "Investigation into the Suzuki-Miyaura coupling aiming at multikilogram synthesis of E2040 using (0-cyanophenyl)boronic esters", Y. Urawa et al., J. Organometallic Chemistry, 653 (2002), 269-278). Also in the present invention, the same effect can be obtained by the addition of water. Water may be used in an amount of 1 to 20 moles, suitably 1 to 10 moles, more suitably 3 to 5 moles, per mole of the compound (I).

When made into a salt, 3-(2-cyanophenyl)-5-(2-pyridyl)-1-phenyl-1,2-dihydropyridine-2-one (III) may be stably isolated as substantially colorless crystals.

Preferable examples of the "salt" are hydrohalogenic acid salts such as hydrofluoride, hydrochloride, hydrobromide, hydroiodide, etc.; inorganic acid salts such as sulfate, nitrate, perchlorate, phosphate, etc.; and organic sulfonates such as methanesulfonate, trifluoromethane-sulfonate, ethanesulfonate, benzenesulfonate, toluenesulfonate, camphorsulfonate, etc. More preferable examples thereof are hydrohalogenic acid salts such as hydrofluoride, hydrochloride, hydrobromide, hydroiodide, etc.; and inorganic acid salts such as sulfate, nitrate, perchlorate, phosphate, etc. The most preferable examples thereof are hydrochloride, hydrofluoride and carbonate.

The crystal forms and production method thereof of the present invention is explained below in detail.

#### Crystal Forms of the Present Invention

The crystal forms of the present invention are crystal forms of the hydrate of Compound (A) with the characteristics described below. Although the respective measurement conditions for powder X-ray diffraction patterns and infrared absorption spectra (KBr method) are not particularly limited, the measurement should preferably be conducted under the measurement conditions for the powder X-ray diffraction patterns and the infrared absorption spectra (KBr method) as will be described below.

(1) A crystal having a diffraction peak at a diffraction angle ( $2\theta \pm 0.2^\circ$ ) of  $8.7^\circ$  in a powder X-ray diffraction;

(2) A crystal having a diffraction peak at a diffraction angle ( $2\theta \pm 0.2^\circ$ ) of  $12.5^\circ$  in a powder X-ray diffraction;

(3) A crystal having diffraction peaks at diffraction angles ( $2\theta \pm 0.2^\circ$ ) of  $8.7^\circ$  and  $12.5^\circ$  in a powder X-ray diffraction;

(4) A crystal having diffraction peaks at the diffraction angles ( $2\theta \pm 0.2^\circ$ ) shown in FIG. 4 or Table 5 below in a powder X-ray diffraction;

(5) A crystal having an absorption peak at a wavelength of  $1588 \pm 1 \text{ cm}^{-1}$  in an infrared absorption spectrum (KBr method);

(6) A crystal having absorption peaks at wavelengths of  $1588 \pm 1 \text{ cm}^{-1}$  and  $751 \pm 1 \text{ cm}^{-1}$  in an infrared absorption spectrum (KBr method); and

(7) A crystal having absorption peaks at the wavelengths ( $\text{cm}^{-1}$ ) shown in FIG. 1 or Table 2 below in an infrared absorption spectrum (KBr method).

These characteristic peaks in the powder X-ray diffraction are not observable in the crystal form obtained by the production process disclosed in WO 01/96308 (see Reference Example 1, Table 4 and FIG. 3 as described below).

As for a diffraction angle ( $2\theta$ ) in the powder X-ray diffraction analysis, errors in the diffraction angle, generally, may occur within the range of  $\pm 0.2^\circ$ . It is, therefore, to be understood that the values of the diffraction angles may include numerals on the order of  $\pm 0.2^\circ$ . Accordingly, this invention encompasses not only crystal form having completely matching diffraction angles of the peaks in powder X-ray diffraction, but also crystal form having matching diffraction angles of the peaks within the errors of about  $\pm 0.2^\circ$ .

#### Crystal of Hydrate

The term "having a diffraction peak at a diffraction angle ( $2\theta \pm 0.2^\circ$ ) of  $8.7^\circ$ " means "having a diffraction peak at a diffraction angle ( $2\theta$ ) of  $8.5^\circ$  to  $8.9^\circ$ ." The term "having a diffraction peak at a diffraction angle ( $2\theta \pm 0.2^\circ$ ) of  $12.5^\circ$ " means "having a diffraction peak at a diffraction angle ( $2\theta$ ) of  $12.3^\circ$  to  $12.7^\circ$ ."

The term "having an absorption peak at a wavenumber of  $1588\pm 1\text{ cm}^{-1}$ " means "having an absorption peak at a wavenumber of  $1587$  to  $1589\text{ cm}^{-1}$ ."

The term "having absorption peaks at wavenumbers of  $1588\pm 1\text{ cm}^{-1}$  and  $751\pm 1\text{ cm}^{-1}$ " means "having absorption peaks at wavenumbers of  $1587$  to  $1589\text{ cm}^{-1}$  and of  $750$  to  $752\text{ cm}^{-1}$ ."

The term "having a peak at chemical shifts of around  $146.7\text{ ppm}$ " means "having a peak substantially equivalent to  $146.7\text{ ppm}$  when a  $^{13}\text{C}$  Solid State NMR spectrum is measured under normal conditions or under the conditions substantially the same as those described in present specification." The term "having a peak at chemical shifts of around  $123.3\text{ ppm}$ " means "having a peak substantially equivalent to  $123.3\text{ ppm}$  when a  $^{13}\text{C}$  Solid State NMR spectrum is measured under normal conditions or under the conditions substantially the same as those described in the present specification."

#### Anhydrous Crystal Form I

The term "having a diffraction peak at a diffraction angle ( $2\theta\pm 0.2^\circ$ ) of  $10.3^\circ$ " means "having a diffraction peak at a diffraction angle ( $2\theta$ ) of  $10.1^\circ$  to  $10.5^\circ$ ." The term "having a diffraction peak at a diffraction angle ( $2\theta\pm 0.2^\circ$ ) of  $19.1^\circ$ " means "having a diffraction peak at a diffraction angle ( $2\theta$ ) of  $18.9^\circ$  to  $19.3^\circ$ ."

The term "having a peak at chemical shifts of around  $149.0\text{ ppm}$ " means "having a peak substantially equivalent to  $149.0\text{ ppm}$  when a  $^{13}\text{C}$  Solid State NMR spectrum is measured under normal conditions or under the conditions substantially the same as those described in the specification." The term "having a peak at chemical shifts of around  $125.6\text{ ppm}$ " means "having a peak substantially equivalent to  $125.6\text{ ppm}$  when a  $^{13}\text{C}$  Solid State NMR spectrum is measured under normal conditions or under the conditions substantially the same as those described in the present specification."

#### Anhydrous Crystal Form V

The term "having a diffraction peak at a diffraction angle ( $2\theta\pm 0.2^\circ$ ) of  $16.7^\circ$ " means "having a diffraction peak at a diffraction angle ( $2\theta$ ) of  $16.5^\circ$  to  $16.9^\circ$ ." The term "having a diffraction peak at a diffraction angle ( $2\theta\pm 0.2^\circ$ ) of  $12.9^\circ$ " means "having a diffraction peak at a diffraction angle ( $2\theta$ ) of  $12.7^\circ$  to  $13.1^\circ$ ." The term "having a diffraction peak at a diffraction angle ( $2\theta\pm 0.2^\circ$ ) of  $24.9^\circ$ " means "having a diffraction peak at a diffraction angle ( $2\theta$ ) of  $24.7^\circ$  to  $25.1^\circ$ ."

The term "having an absorption peak at a wavenumber of  $1658\pm 1\text{ cm}^{-1}$ " means "having an absorption peak at a wavenumber of  $1657$  to  $1659\text{ cm}^{-1}$ ."

The term "having an absorption peak at a wavenumber of  $501\pm 1\text{ cm}^{-1}$ " means "having an absorption peak at a wavenumber of  $500$  to  $502\text{ cm}^{-1}$ ."

The term "having a peak at chemical shifts of around  $145.9\text{ ppm}$ " means "having peak substantially equivalent to  $145.9\text{ ppm}$  when a  $^{13}\text{C}$  Solid State NMR spectrum is measured under normal conditions or under the conditions substantially the same as those described in the specification." The term "having a peak at chemical shifts of around  $137.7\text{ ppm}$ " means "having a peak substantially equivalent to  $137.7\text{ ppm}$  when a  $^{13}\text{C}$  Solid State NMR spectrum is measured under normal conditions or under the conditions substantially the same as those described in the present specification."

#### Anhydrous Crystal Form III

The term "having a diffraction peak at a diffraction angle ( $2\theta\pm 0.2^\circ$ ) of  $23.7^\circ$ " means "having a diffraction peak at a diffraction angle ( $2\theta$ ) of  $23.5^\circ$  to  $23.9^\circ$ ." The term "having a diffraction peak at a diffraction angle ( $2\theta\pm 0.2^\circ$ ) of  $25.0^\circ$ " means "having a diffraction peak at a diffraction angle ( $2\theta$ ) of  $24.8^\circ$  to  $25.2^\circ$ ." The term "having a diffraction peak at a diffraction angle ( $2\theta\pm 0.2^\circ$ ) of  $5.7^\circ$ " means "having a diffraction

peak at a diffraction angle ( $2\theta$ ) of  $5.5^\circ$  to  $5.9^\circ$ ." The term "having a diffraction peak at a diffraction angle ( $2\theta\pm 0.2^\circ$ ) of  $9.5^\circ$ " means "having a diffraction peak at a diffraction angle ( $2\theta$ ) of  $9.3^\circ$  to  $9.7^\circ$ ."

The term "an alkyl ketone solvent" means an organic solvent of dialkyl ketone such as acetone and ethyl methyl ketone, and preferably acetone.

The term "an alcoholic solvent" means an organic solvent of  $\text{C}_{1-6}$  alcohol, such as methanol, ethanol, 1-propanol and 2-propanol, and preferably methanol or 1-propanol.

The term "under reduced pressure" is not particularly limited insofar as it is  $760\text{ mmHg}$  or less; and it is preferably from  $760$  to  $0.1\text{ mmHg}$ , more preferably from  $50$  to  $0.1\text{ mmHg}$ , and most preferably from  $30$  to  $5\text{ mmHg}$ .

The production methods of the crystal forms of the present invention are explained below in detail.

#### General Production Process for Crystal of Hydrate

The crystal of hydrate of the present invention can be stably produced on an industrial scale by preparing Compound (A) according to Example 7 in WO 01/96308 or Example 1 as described herein below, dissolving Compound (A) in a specific solvent by heating, and crystallizing it by cooling at stirring.

Compound (A) to be used in the crystallization may be any form of hydrate, anhydrous, amorphous and crystalline forms which includes plural crystal polymorphs, and may even be a mixture of the foregoing.

The solvents to be used in the crystallization include one member or a mixed solvent of two members selected from the group consisting of an alcoholic solvent, an alkyl ketone solvent, and water. The solvent is preferably a mixed solvent of acetone and water.

When the mixed solvent of acetone and water is used, its mixing ratio (volume ratio) is preferably from  $37:3$  to  $24:16$ , more preferably from  $9:1$  to  $7:3$ , and further more preferably about  $8:2$ . The most preferred is a mixed solvent formed by dissolving the crystals in a mixed solvent of acetone and water ( $9:1$ ) and thereafter adding water to the mixed solvent to prepare a solution of acetone and water ( $8:2$ ).

The amount of the solvent used may appropriately be selected between the lower limit and the upper limit, the lower limit being an amount to dissolve Compound (A) by heating and the upper limit being an amount so as not to significantly reduce the yield of the crystals. The amount of crystallization solvent is preferably from  $10$ - to  $50$ -fold (v/w) as a volume ratio based on the weight of Compound (A), more preferably from  $30$ - to  $50$ -fold (v/w). Further preferably, the amount is about  $40$ -fold (v/w) if acetone-water ( $9:1$ ) is used; and it is about  $45$ -fold (v/w) if acetone-water ( $8:2$ ) is used.

The temperature at which Compound (A) is dissolved by heating may appropriately be selected as the temperature to dissolve Compound (A1) depending on the solvent. The temperature is preferably from the reflux temperature of the crystallization solvent to  $50^\circ\text{C}$ ., more preferably from  $65$  to  $55^\circ\text{C}$ .

Any change in the cooling rate during crystallization may produce crystals with different forms (polymorphisms). It is, therefore, desired that the crystallization be performed by suitably adjusting the cooling rate in consideration of possible effects on the quality and the particle size of the crystals or the like. Cooling is preferably performed at a rate of  $40$  to  $5^\circ\text{C}$ . per hour and more preferably at a rate of  $25$  to  $15^\circ\text{C}$ . per hour.

The final crystallization temperature may also appropriately be selected in consideration of the yield and the quality of the crystals, or the like; and it is preferably from  $10$  to  $-25^\circ\text{C}$ .

## 21

In the crystallization, seed crystals may or may not be added, which comprise a small amount of crystals of 3-(2-cyanophenyl)-5-(2-pyridyl)-1-phenyl-1,2-dihydropyridin-2-one hydrate. There is no particular limitation to the temperature at which the seed crystals are added. The temperature is preferably 60° C. or less, more preferably from 55 to 0° C., further more preferably from 55 to 35° C., and most preferably about 40° C.

The precipitated crystals can be separated by usual filtration, washed with solvent if necessary, and then dried to produce the desired crystals. The solvent for use in washing the crystals is common to the crystallization solvent, and it is preferably a mixed solvent of acetone-water (9:1 to 1:1), more preferably a mixed solvent of acetone-water (about 1:1).

## Drying Method for Crystals

The crystals separated by the filtration can be dried by allowing them to stand in the atmosphere, where appropriate, or by heating.

The time during which the residual solvent is removed below the prescribed amount may appropriately be selected as the drying time, depending on the production quantity, the drying apparatus, the drying temperature or the like. Drying can be performed either under aeration or under reduced pressure. The level of pressure reduction may appropriately be selected, depending on the production quantity, the drying apparatus, the drying temperature or the like. The obtained crystals may be allowed to stand in the atmosphere after drying if necessary.

The crystals produced by the above-mentioned process comprise a homogeneous crystal form. Being provided with the preferable properties such that it is stable, has no tendency to readily transform into other crystal or amorphous forms, and is not hygroscopic, these crystals are suited to formulation.

The use of Compound (A) as a therapeutic agent for neurodegenerative diseases or others is fully disclosed in WO 01/96308. The crystal forms of the invention can be used as the active ingredient in the therapeutic agent for neurodegenerative diseases or others. The entire disclosure of WO 01/96308 is thus hereby incorporated by reference.

When Compound (A) of the present invention is to be used as a medicament, it is normally compounded with suitable pharmaceutical ingredients to prepare pharmaceutical products for use. Notwithstanding, the use of a drug substance form of the compound of the invention as a medicament should not be negated.

The pharmaceutical ingredients may include excipients, binders, lubricants, disintegrating agents, coloring agents, taste correctives, emulsifiers, surfactants, dissolving aids, suspending agents, isotonicizing agents, buffering agents, preservatives, antioxidants, stabilizers, absorption enhancers, and the like, all of which are generally used in medicaments. If desired, these agents may be combined for use.

The excipients may include, for example, lactose, white soft sugar, glucose, corn starch, mannitol, sorbitol, starch, alpha starch, dextrin, crystalline cellulose, light silicic anhydride, aluminum silicate, calcium silicate, magnesium aluminometasilicate, calcium hydrogenphosphate, and the like.

The binders may include, for example, polyvinyl alcohol, methylcellulose, ethylcellulose, gum Arabic, tragacanth, gelatin, shellac, hydroxypropylmethylcellulose, hydroxypropylcellulose, carboxymethylcellulose sodium, polyvinylpyrrolidone, macrogol, and the like.

The lubricants may include, for example, magnesium stearate, calcium stearate, sodium stearyl fumarate, talc, polyethylene glycol, colloidal silica, and the like.

## 22

The disintegrating agents may include, for example, crystalline cellulose, agar, gelatin, calcium carbonate, sodium hydrogencarbonate, calcium citrate, dextrin, pectin, low-substituted hydroxypropylcellulose, carboxymethylcellulose, carboxymethylcellulose calcium, croscarmellose sodium, carboxymethyl starch, carboxymethyl starch sodium, and the like.

The coloring agents may include iron sesquioxide, yellow iron sesquioxide, carmine, caramel, beta-carotene, titanium oxide, talc, riboflavin sodium phosphate, yellow aluminum lake, and the like, which have been approved as additives for medicaments.

The taste correctives agents may include cocoa powder, menthol, aromatic powder, mentha oil, borneol, powdered cinnamon bark, and the like.

The emulsifiers or the surfactants may include stearyl triethanolamine, sodium lauryl sulfate, lauryl aminopropionic acid, lecithin, glycerin monostearate, sucrose fatty acid ester, glycerin fatty acid ester, and the like.

The dissolving aids may include polyethylene glycol, propylene glycol, benzyl benzoate, ethanol, cholesterol, triethanolamine, sodium carbonate, sodium citrate, Polysorbate 80, nicotinamide, and the like.

The suspending agents may include, in addition to the surfactants, hydrophilic polymers such as polyvinyl alcohol, polyvinylpyrrolidone, methylcellulose, hydroxymethylcellulose, hydroxyethylcellulose, and hydroxypropylcellulose.

The isotonicizing agents may include glucose, sodium chloride, mannitol sorbitol and the like.

The buffering agents may include the buffers of phosphate, acetate, carbonate, citrate and the like.

The preservatives may include methylparaben, propylparaben, chlorobutanol, benzyl alcohol, phenethyl alcohol, dehydroacetic acid, sorbic acid and the like.

The antioxidants may include sulfite, ascorbic acid, alpha-tocopherol and the like.

The stabilizers may include those generally used in medicaments.

The absorption enhancers may include those generally used in medicaments.

The pharmaceutical products described above may include: oral agents such as tablets, powders, granules, capsules, syrups, troches, and inhalations; external preparations such as suppositories, ointments, ophthalmic ointments, tapes, ophthalmic solutions, nasal drops, ear drops, poultices, and lotions; and injections.

The oral agents may appropriately be combined with the auxiliaries described above to form preparations. In addition, the surfaces of the agents may be coated if necessary.

The external preparations may appropriately be combined with the auxiliaries, in particular, excipients, binders, taste correctives, emulsifiers, surfactants, dissolving aids, suspending agents, isotonicizing agents, preservatives, antioxidants, stabilizers, or absorption enhancers to form the preparations.

The injections may appropriately be combined with the auxiliaries, in particular, emulsifiers, surfactants, dissolving aids, suspending agents, isotonicizing agents, preservatives, antioxidants, stabilizers, or absorption enhancers to form the preparations.

When Compound (A) of the present invention is to be used as a medicament, its dosage level may differ depending on the symptoms, ages or others. Compound (A) is normally given in a single administration or in divided administrations 2 to 6 times daily at the following doses: from 0.05 to 10 mg (preferably from 0.1 to 5 mg) in the case of an oral agent; from 0.01 to 10 mg (preferably from 0.05 to 5 mg) in the case of an

## 23

external preparation; and 0.01 to 5 mg in the case of an injection. Here, the actual amounts to be administered are indicated with respect to the oral agent and the injection, while the amount to be absorbed by the body is indicated with respect to the external preparation.

The preparations for therapeutic or prophylactic use in humans containing crystal of 3-(2-cyanophenyl)-5-(2-pyridyl)-1-phenyl-1,2-dihydropyridin-2-one hydrate (Compound (A)) according to the invention may be obtained by the general methods that are accepted in manufacturing pharmacy. The specific formulation examples of preparations are shown below.

The compound of the present invention [i.e., 3-(2-cyanophenyl)-5-(2-pyridyl)-1-phenyl-1,2-dihydropyridin-2-one], lactose, low-substituted hydroxypropylcellulose were blended. Polyvinylpyrrolidone dissolved in an appropriate amount of purified water was then used to wet granulate the blend. These granulates were dried and then size-controlled. Low-substituted Hydroxypropylcellulose and magnesium stearate were blended to the resulting granulates, after which they was tableted. The obtained tablets were film-coated with an aqueous solution of the coating base (a mixture of hydroxypropylmethylcellulose, talc, Macrogol 6000, titanium oxide and yellow iron sesquioxide). The amounts of the respective materials to be used per tablet are shown in Table 1 below.

TABLE 1

Material	Purpose	0.5 mg tablet	1.0 mg tablet	2.0 mg tablet
compound of the invention* <sup>1</sup>	principle agent	0.5 mg	1.0 mg	2.0 mg
lactose	excipient	80.0 mg	79.5 mg	78.5 mg
low-substituted hydroxypropylcellulose	disintegrator	9.0 mg	9.0 mg	9.0 mg
polyvinylpyrrolidone	binder	5.0 mg	5.0 mg	5.0 mg
low-substituted hydroxypropylcellulose	disintegrator	5.0 mg	5.0 mg	5.0 mg
magnesium stearate	lubricant	0.5 mg	0.5 mg	0.5 mg
purified water	solvent	q.s.	q.s.	q.s.
coating base* <sup>2</sup>	coating agent	5.0 mg	5.0 mg	5.0 mg
purified water	solvent	q.s.	q.s.	q.s.
total		105 mg	105 mg	105 mg

\*<sup>1</sup>3-(2-Cyanophenyl)-5-(2-pyridyl)-1-phenyl-1,2-dihydropyridin-2-one (Hydrate)

\*<sup>2</sup>Mixture of hydroxypropylmethylcellulose, talc, Macrogol 6000, titanium oxide and yellow iron sesquioxide

The present invention is explained below in further detail with working examples but they are merely for illustration and the production method of the present invention is not limited in any case by the following specific examples. Those skilled in the art may conduct the present invention to maximum by making various modifications to not only the follow-

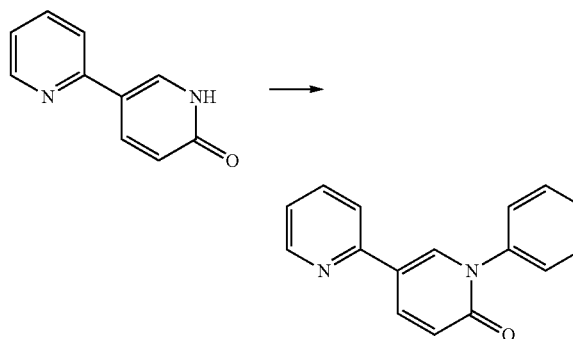
## 24

ing working examples but also the claims in the present specification, and these modifications are included in the claims in the present specification.

## EXAMPLE 1

Synthesis of 3-(2-cyanophenyl)-5-(2-pyridyl)-1-phenyl-1,2-dihydropyridine-2-one

(1) Synthesis of 5-(2-pyridyl)-1-phenyl-1,2-dihydropyridine-2-one



After the inner atmosphere of a reactor was replaced with nitrogen, a mixture of 5-(2-pyridyl)-1,2-dihydropyridine-2-one (WO2004/009553) (7.33 kg), triphenylboroxine (9.0 kg), copper acetate (anhydrous) (0.80 kg), water (0.50 kg), pyridine (7.1 kg) and N,N -dimethylformamide (66.7 kg) was stirred for 1 hour in the reactor at an internal temperature of 28° C.

While introducing air adjusted to an oxygen concentration of 9% with nitrogen into the reactor at a rate of 30 L/min, the reaction mixture was stirred for 16 hours at 39° C. to 40° C. (internal temperature) to obtain a reaction mixture 1A.

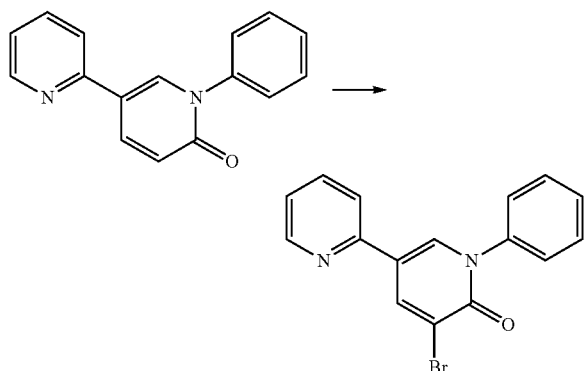
Water (191 kg) and 25% aqueous ammonia (85.8 kg) were placed in another reactor and cooled to 8.7° C. with cold water. Then, the above-mentioned reaction mixture 1A was added thereto over a period of 3 minutes. The resulting reaction mixture was stirred for 4 hours while being cooled with cold water. The precipitate in the reaction mixture was collected by filtration by the use of a centrifuge and washed with 65 kg of water.

The precipitate, water (97 kg) and 25% aqueous ammonia (43.5 kg) were placed in a reactor and stirred for 1 hour while being kept warm with warm water at 25° C. The precipitate in the reaction mixture was collected by filtration by the use of a centrifuge, washed with 32.6 kg of water and then dried under reduced pressure (60° C., 18 hours) to obtain 9.6 kg of 5-(2-pyridyl)-1-phenyl-1,2-dihydropyridine-2-one.

25

<sup>1</sup>H NMR (400 MHz, DMSO-d<sub>6</sub>) δ 8.61-8.50 (m, 1H), 8.36 (d, 1H), 8.29 (dd, 1H), 7.90 (d, 1H), 7.80 (ddd, 1H), 7.56-7.45 (m, 5H), 7.27 (dd, 1H), 6.62 (d, 1H).

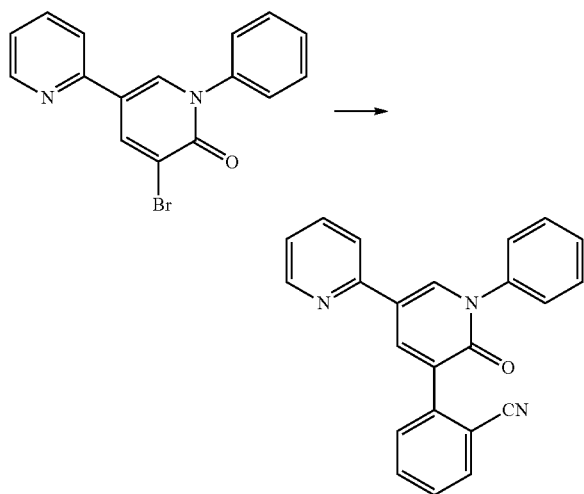
(2) Synthesis of 3-bromo-5-(2-pyridyl)-1-phenyl-1,2-dihydropyridine-2-one



5-(2-Pyridyl)-1-phenyl-1,2-dihydropyridine-2-one (200 g), N-bromosuccinimide (157.7 g) and ethyl acetate (4 L) were placed in a 10-L reactor, and the reaction mixture was stirred at 30° C. (external temperature) in a nitrogen stream for 9 hours and 20 minutes. A 3% aqueous hydrosulfite solution (2 L) and toluene (2 L) were added to the reaction mixture, followed by stirring at 55° C. (external temperature) for 30 minutes. After completion of the reaction, the aqueous layer (the lower layer) in the reaction mixture was separated. Then, the organic layer was washed four times with water (2 L), and the organic solvent was removed under reduced pressure with stirring.

Thereafter, 1,2-dimethoxyethane (4 L) was added to the residue and the resulting mixture was concentrated under reduced pressure to obtain crude 3-bromo-5-(2-pyridyl)-1-phenyl-1,2-dihydropyridine-2-one.

(3) Synthesis of 3-(2-cyanophenyl)-5-(2-pyridyl)-1-phenyl-1,2-dihydropyridine-2-one



26

2-(1,3,2-Dioxaborinan-2-yl)benzonitrile (214.9 g), palladium acetate (3.44 g), triphenylphosphine (16.07 g), cuprous iodide (7.29 g), 1,2-dimethoxyethane (3.1 L) and potassium carbonate (158.8 g) were placed in a reactor containing the whole of the crude 3-bromo-5-(2-pyridyl)-1-phenyl-1,2-dihydropyridine-2-one obtained as concentration residue in the above item (2), and the resulting mixture was stirred with heating at 70° C. (external temperature) for 30 minutes under a nitrogen atmosphere and then stirred with heating under reflux for 4 hours.

Thereafter, ethyl acetate (2.5 L) was added to the reaction mixture at 70° C. (external temperature) and stirred for 10 minutes. The resulting reaction mixture was filtered and the precipitate was washed with ethyl acetate (2.5 L). The whole of the filtrate thus obtained was transferred into a reactor and 12.5% aqueous ammonia (5 L) was added thereto, followed by stirring at 60° C. (external temperature) for 53 minutes. The lower layer (the aqueous layer) in the reaction mixture was separated. A 5% aqueous sodium chloride solution (2.5 L) and 25% aqueous ammonia (2.5 L) were added to the remaining organic layer and stirred. Thereafter, the lower layer (the aqueous layer) was separated and a 5% aqueous sodium chloride solution (5 L) was added to the remaining organic layer and stirred, and then the lower layer (the aqueous layer) was separated. The remaining organic layer was concentrated under reduced pressure, followed by adding thereto 4 L of acetone, and the resulting mixture was concentrated under reduced pressure.

Acetone (7.2 L) and water (0.8 L) were added to the residue and the resulting mixture was stirred at 60° C. (external temperature) for 1 hour and 10 minutes to effect dissolution. The resulting solution was cooled with stirring at 38° C. (external temperature) for 18 minutes. To the reaction mixture was added 1 g of seed crystals (crystals of hydrate of 3-(2-cyanophenyl)-5-(2-pyridyl)-1-phenyl-1,2-dihydropyridine-2-one) at an internal temperature of 40° C., and the resulting mixture was stirred at 35° C. (external temperature) for 30 minutes. Thereafter, the reaction mixture was stirred while lowering the external temperature by 5° C. at intervals of 30 minutes. At an external temperature of 10° C., the reaction mixture was stirred for 17 hours.

Water (2.29 L) was added dropwise to the reaction mixture with stirring over a period of 3 hours and 10 minutes. After completion of the dropwise addition, the resulting mixture was stirred for another 1 hour and 20 minutes. The reaction mixture was filtered and the precipitate was washed with 2 L of 50% acetone-water to obtain 3-(2-cyanophenyl)-5-(2-pyridyl)-1-phenyl-1,2-dihydropyridine-2-one (526.28 g) as a wet cake (dry weight: 168.3 g).

(4) Conversion of 3-(2-cyanophenyl)-5-(2-pyridyl)-1-phenyl-1,2-dihydropyridine-2-one in the Wet Cake to Dried Weight

The obtained wet cake (4.378 g) was weighed out and dried under reduced pressure at 50° C. for 4 hours to give 1.4005 g of a dried powder.

$$\text{Converted value as dried weight} = (1.4005/4.378) \times 526.28 = 168.3 \text{ g}$$

(5) Determination of Acetone and Water Weight Contents in the Wet Cake of 3-(2-cyanophenyl)-5-(2-pyridyl)-1-phenyl-1,2-dihydropyridine-2-one

Gas chromatographic analysis of the obtained wet cake under the conditions described below ascertained that the wet cake obtained in (3) above contained 168 mL of acetone and 186 mL of water.

27

Gas Chromatographic Analysis Conditions:

Column: DB-WAX (30 m×0.53 mm, 1 μm); Detector: TCD;

Oven temp.: 60° C. (8 min), 60-180° C. (70° C./min), 180° C. (5 min);

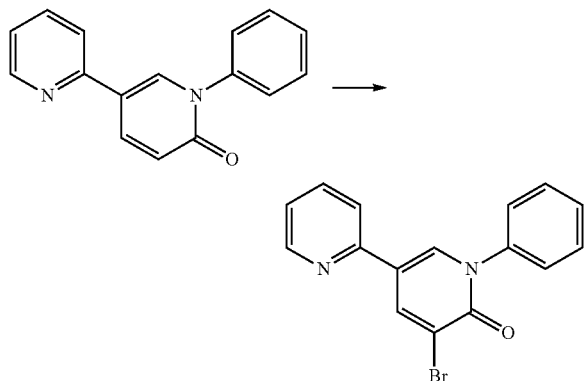
Detector temp.: 210° C.; inlet temp.: 150° C.; Column flow: 5.0 mL/min; split ratio: (1:4);

Injection vol.: 2 μL

## EXAMPLE 2

Synthesis of 3-(2-cyanophenyl)-5-(2-pyridyl)-1-phenyl-1,2-dihydropyridine-2-one

(1) Synthesis of 3-bromo-5-(2-pyridyl)-1-phenyl-1,2-dihydropyridine-2-one

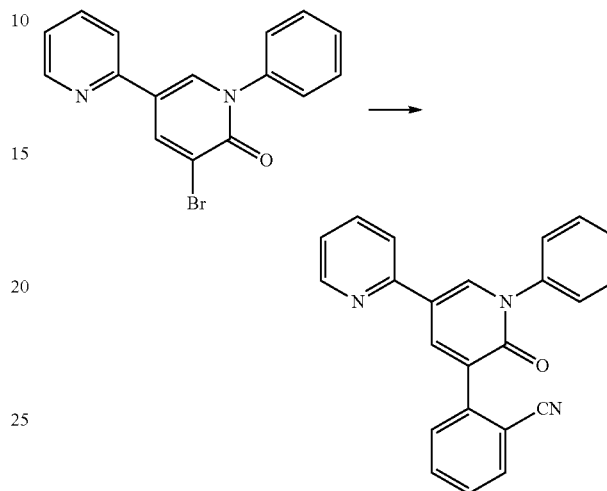


5-(2-Pyridyl)-1-phenyl-1,2-dihydropyridine-2-one (300 g), N-bromosuccinimide (236.5 g) and N,N-dimethylformamide (1.8 L) were placed in a 10-L reactor, and the reaction mixture was stirred at 30° C. (external temperature) in a nitrogen stream for 3 hours and 15 minutes. 2-Propanol (4.2 L) was added dropwise to the reaction mixture over a period of 9 minutes, followed by adding thereto water (2.1 L) over a period of 7 minutes. The resulting mixture was heated at 85° C. (external temperature) with stirring. After confirming the dissolution of the contents, the resulting solution was stirred at an external temperature of 55° C. for 1 hour. Thereafter, the solution was stirred at 40° C. (external temperature) for another 22 minutes, at 30° C. (external temperature) for further another 23 minutes, and then at 10° C. (external temperature) for still another 15 hours and 15 minutes. The reaction mixture was filtered and the precipitate was washed with 50% 2-propanol-water (2.4 L) and then dried under reduced pressure (60° C., 6 hours) to obtain 341.45 g of 3-bromo-5-(2-pyridyl)-1-phenyl-1,2-dihydropyridine-2-one. Yield: 86.4%.

28

<sup>1</sup>H NMR (400 MHz, CDCl<sub>3</sub>) δ 8.59-8.56 (m, 1H), 8.50 (d, 1H), 8.18 (d, 1H), 7.72 (td, 1H), 7.53-7.41 (m, 6H), 7.20 (ddd, 1H).

(2) Synthesis of 3-(2-cyanophenyl)-5-(2-pyridyl)-1-phenyl-1,2-dihydropyridine-2-one



3-Bromo-5-(2-pyridyl)-1-phenyl-1,2-dihydropyridine-2-one (100 g), 2-(1,3,2-dioxaborinan-2-yl)benzonitrile (85.7 g), palladium acetate (1.37 g), triphenylphosphine (6.4 g), cuprous iodide (2.91 g), 1,2-dimethoxyethane (1.25 L) and potassium carbonate (63.4 g) were placed in a 3-L reactor, and pressure reduction and the replacement of the air in the reaction system with nitrogen by repressurization with nitrogen were carried out 10 times. The reaction mixture was stirred with heating (in an oil bath at 100° C.) under a nitrogen atmosphere for 3 hours and 40 minutes.

Thereafter, ethyl acetate (750 mL) was added to the reaction mixture and the resulting mixture was filtered. The precipitate was washed with ethyl acetate (750 mL). To the filtrate thus obtained were added 750 mL of water and 25% aqueous ammonia (250 mL), and the resulting mixture was stirred at 60° C. (external temperature) for 30 minutes. The lower layer (the aqueous layer) in the reaction mixture was separated. A 2.5% aqueous sodium chloride solution (370 mL), 25% aqueous ammonia (130 mL) and 1,2-dimethoxyethane (500 mL) were added to the remaining organic layer, followed by stirring at 60° C. (external temperature) for 10 minutes. Thereafter, the lower layer (the aqueous layer) was separated and a 2.5% aqueous sodium chloride solution (370 mL), 25% aqueous ammonia (130 mL) and 1,2-dimethoxyethane (200 mL) were added to the remaining organic layer and stirred for 10 minutes, and then the lower layer (the aqueous layer) was separated. A 2.5% aqueous sodium chloride solution (500 mL) and 1,2-dimethoxyethane (200 mL) were added to the remaining organic layer, followed by stirring at 60° C. (external temperature) for 10 minutes. Thereafter, the lower layer (the aqueous layer) was separated. The remaining organic layer was concentrated under reduced pressure (external temperature: 65° C.), followed by adding thereto 2 L of acetone, and the resulting mixture was concentrated under reduced pressure (external temperature: 60° C.). Acetone (2.88 L) and water (320 mL) were added to the residue and the resulting mixture was stirred at 55° C. (external temperature) for 1 hour and 10 minutes to effect dissolu-

29

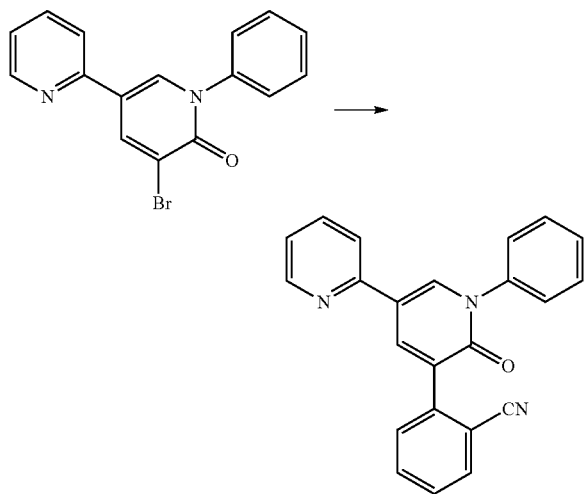
tion. The resulting solution was cooled with stirring at 38° C. (external temperature) for 38 minutes. To the reaction mixture was added 500 mg of seed crystals (crystals of hydrate of 3-(2-cyanophenyl)-5-(2-pyridyl)-1-phenyl-1,2-dihydropyridine-2-one) at an internal temperature of 40° C., and the resulting mixture was stirred for 1 hour at an external temperature changed to 30° C. This mixture was stirred for 1 hour at an external temperature changed to 20° C. and then stirred for 1 hour and 20 minutes at an external temperature of 8° C.

Water (915 mL) was added dropwise to the reaction mixture with stirring over a period of 2 hours and 50 minutes. After completion of the dropwise addition, the resulting mixture was stirred for another 14 hours. The reaction mixture was filtered and the precipitate was washed with 500 mL of 50% acetone-water and then 500 mL of water to obtain 3-(2-cyanophenyl)-5-(2-pyridyl)-1-phenyl-1,2-dihydropyridine-2-one (251.5 g) as a wet cake (dry weight: 83.3 g).

<sup>1</sup>H NMR (400 MHz, DMSO-d<sub>6</sub>) δ 8.61-8.57 (m, 1H), 8.52 (d, 1H), 8.47 (d, 1H), 8.00 (d, 1H), 7.92 (d, 1H), 7.83 (td, 1H), 7.78 (t, 1H), 7.74-7.70 (d-like, 1H), 7.61-7.48 (m, 6H), 7.29 (dd, 1H).

## EXAMPLE 3

Synthesis of 3-(2-cyanophenyl)-5-(2-pyridyl)-1-phenyl-1,2-dihydropyridine-2-one



3-Bromo-5-(2-pyridyl)-1-phenyl-1,2-dihydropyridine-2-one (188 g), 2-(1,3,2-dioxaborinan-2-yl)benzonitrile (161.2 g), palladium acetate (2.58 g), triphenylphosphine (12.07 g), 1,2-dimethoxyethane (2.82 L) and ion-exchanged water (41.4 mL) were placed in a 5-L reactor, and pressure reduction and the replacement of the air in the reaction system with nitrogen by repressurization with nitrogen were carried out 5 times with stiffling. Potassium carbonate (119.14 g) was added to the reaction mixture and pressure reduction and the replacement of the air in the resulting mixture with nitrogen by repressurization with nitrogen were carried out 5 times. Then, the reaction mixture was stirred with heating (in an oil bath at 95° C.) under reflux in a nitrogen atmosphere for 1 hour and 49 minutes.

Thereafter, the oil bath was removed and ethyl acetate (800 mL) was added to the reaction mixture at 65.4° C. (internal temperature). The resulting mixture was filtered and the precipitate was washed with ethyl acetate (2.4 L). The filtrate

30

(5.28 kg) thus obtained was divided into halves (2.64 kgx2) and each half was transferred into a 5-L reactor. Trimercaptotriazine (3.05 g) and ethyl acetate (380 mL) were placed in each of the reactors and the reaction mixture was stirred at 50° C. (the external temperature in an oil bath) for 13 hours and 10 minutes. The two solutions thus obtained were filtered in succession by the use of Celite (94 g) previously rinsed with methanol (1 L) and ethyl acetate (1 L) and the precipitate was rinsed with a 4:3 mixture (1.35 L) of ethyl acetate and 1,2-dimethoxyethane. The filtrate thus obtained was transferred into a 20-L separator and hydrochloric acid prepared from concentrated hydrochloric acid (700 mL) and ion-exchanged water (4.2 L) was added to the filtrate. After stirring at 37.6° C. (internal temperature) for 8 minutes, the aqueous layer (the lower layer) was separated. Then, 2N-hydrochloric acid (3.8 L) was added to the organic layer, followed by stirring at 39.3° C. (internal temperature) for 8 minutes, and the aqueous layer (the lower layer) was separated. Ethyl acetate (3 L) was added to the combined aqueous layer and stirred for 8 minutes, and then ethyl acetate (3 L) was added thereto and stirred for 5 minutes. Thereafter, the aqueous layer (the lower layer) was separated. This aqueous layer was cooled to 20° C. (internal temperature) with stirring in a cold-water bath, and then 25% aqueous ammonia (2.25 L) was added dropwise thereto over a period of 27 minutes with cooling in an ice-water bath. The resulting mixture was stirred for another 3 hours and 26 minutes. The reaction mixture was filtered under reduced pressure and the precipitate was washed with ion-exchanged water (3 L). The washed precipitate was dried by air blowing (60° C., 16 hours and 6 minutes) to obtain 3-(2-cyanophenyl)-5-(2-pyridyl)-1-phenyl-1,2-dihydropyridine-2-one (162.62 g).

<sup>1</sup>H NMR (400 MHz, DMSO-d<sub>6</sub>) δ 8.60-8.57 (m, 1H), 8.53 (d, 1H), 8.47 (d, 1H), 8.00 (d, 1H), 7.92 (d, 1H), 7.83 (td, 1H), 7.78 (t, 1H), 7.72 (d, 1H), 7.61-7.48 (m, 6H), 7.30 (dd, 1H).

## EXAMPLE 4

Crystallization of 3-(2-cyanophenyl)-5-(2-pyridyl)-1-phenyl-1,2-dihydropyridine-2-one (Hydrate Crystal)

A 10 L-flask was charged with 526.28 g of 3-(2-cyanophenyl)-5-(2-pyridyl)-1-phenyl-1,2-dihydropyridine-2-one obtained as the wet cake in (3) of Example 1. Out of an acetone-water prepared from 5890 mL of acetone and 490 mL of water, 5.5 L was added to the flask and heated. Filtration was carried out after dissolution. While the 10-L flask and the filtrated residue were washed with the remaining total of the acetone water, all the filtrate was transferred to a 10-L flask.

The mixture was stirred at an external temperature of 40° C., and after the internal temperature reached 40° C., the external temperature was adjusted to 35° C. Next, 842 mg of 3-(2-cyanophenyl)-5-(2-pyridyl)-1-phenyl-1,2-dihydropyridine-2-one hydrate was added to the mixture. After stirring the mixture for 30 minutes, the external temperature was changed to 30° C., and then to 25° C. after 30 minutes. The external temperature was lowered by 5° C. every 30 minutes thereafter to as low as 15° C. After stiffling the mixture at an external temperature of 15° C. for 30 minutes, the external temperature was further lowered to 8° C. and stirring continued for 1 hour.

To the mixture was added dropwise 842 mL of water at 11° C. (internal temperature) over a period of 1 hour and 10 minutes. One hour after the completion of addition, the external temperature was changed to 0° C. and the mixture was stirred for 40 minutes. The external temperature was then lowered to -20° C. and stirring continued for 15 hours.



## 31

The precipitates in the mixture were collected by filtration. After the precipitates were washed with 1700 mL of 50% acetone-water, they were dried under aeration for 50 minutes. Subsequently, these precipitates were dried with a vibration drier at 40° C. under reduced pressure for 11 hours and were additionally dried at 60° C. for 3 hours.

After the temperature of the drier was cooled to room temperature, the external atmosphere was aspirated into the drier at 950 hpa for 4 hours to give 172.4 g of 3-(2-cyanophenyl)-5-(2-pyridyl)-1-phenyl-1,2-dihydropyridin-2-one (crystal form of the hydrate).

<sup>1</sup>H NMR (400 MHz, DMSO-d<sub>6</sub>): δ 8.61-8.57 (m, 1H), 8.53-8.52 (d-like, 1H), 8.47 (d, 1H), 8.01 (d, 1H), 7.92 (d, 1H), 7.86-7.81 (t-like, 1H), 7.79-7.76 (t-like, 1H), 7.72 (d, 1H), 7.61-7.48 (m, 6H), 7.31-7.28 (m, 1H).

Residual palladium: 15 ppm

## Reference Example 1

### Production Anhydrous Crystals of 3-(2-cyanophenyl)-5-(2-pyridyl)-1-phenyl-1,2-dihydropyridin-2-one (Anhydrous Crystal Form II)

In the same manner as the procedure after reaction work-up that are described in Example 7 in WO 01/96308, the production was carried out below. The synthetic method for 3-(2-cyanophenyl)-5-(2-pyridyl)-1-phenyl-1,2-dihydropyridin-2-one [alternative name: 2-(2-oxo-1-phenyl-5-(pyridin-2-yl)-1,2-dihydropyridin-3-yl)benzonitrile] is described in Example 7 in WO 01/96308 as well as in Example 1 herein above.

Ethyl acetate (400 mL) was added to 3-(2-cyanophenyl)-5-(2-pyridyl)-1-phenyl-1,2-dihydropyridin-2-one (8 g). The mixture was heated at 60° C. in a warm bath. Additional acetate (160 mL) was added to the mixture and the solids were dissolved by heating at 70° C. in the warm bath. After n-hexane (80 mL) was added to this solution, the solvent was evaporated under reduced pressure to give 7.7 g of a pale yellow powder.

<sup>1</sup>H NMR (400 MHz, DMSO-d<sub>6</sub>): δ 8.59-8.57 (m, 1H), 8.53 (d, 1H), 8.47 (d, 1H), 8.01 (d, 1H), 7.92 (d, 1H), 7.83 (ddd, 1H), 7.80-7.76 (m, 1H), 7.73-7.71 (d-like, 1H), 7.61-7.48 (m, 6H), 7.30 (dd, 1H).

## EXAMPLE 5

### Production of Crystals of 3-(2-cyanophenyl)-5-(2-pyridyl)-1-phenyl-1,2-dihydropyridin-2-one hydrate (Hydrate Crystal)

A 500 mL-egg plant vessel was charged with 7 g of 3-(2-cyanophenyl)-5-(2-pyridyl)-1-phenyl-1,2-dihydropyridin-2-one. To the vessel was added 280 mL of 90% acetone-water prepared from 252 mL of acetone and 28 mL of water. The mixture was stirred at heating in a water bath and dissolved under reflux (water bath; 65° C.). After the dissolution was confirmed, the water bath was cooled to 50° C. After addition of 35 mL of water, 140 mg of seed crystals [a small amount of crystals of 3-(2-cyanophenyl)-5-(2-pyridyl)-1-phenyl-1,2-dihydropyridin-2-one hydrate] was added to the vessel at an internal temperature of 50° C. A thermostat bath was used to cool the mixture to -20° C. at a cooling rate of about 35° C./hour. After stifling the mixture at -20° C. for 1 hour, the precipitated solids were collected by filtration and dried under reduced pressure (at an external temperature of 30° C. for 1 hour and then at 60° C. for 2 hours). The obtained dried powders (6.3 g) were transferred to a Petri dish and were allowed to stand in the atmosphere for 17 hours (humidity

## 32

before standing: 55.4%; humidity after overnight standing: 61.6%) to give 6.2 g of crystals of 3-(2-cyanophenyl)-5-(2-pyridyl)-1-phenyl-1,2-dihydropyridin-2-one hydrate.

<sup>1</sup>H NMR (400 MHz, DMSO-d<sub>6</sub>): δ 8.61-8.57 (m, 1H), 8.53 (d, 1H), 8.47 (d, 1H), 8.01 (d, 1H), 7.92 (d, 1H), 7.83 (ddd, 1H), 7.78 (ddd, 1H), 7.73-7.71 (d-like, 1H), 7.61-7.48 (m, 6H), 7.30 (dd, 1H).

Analysis of water content (Karl Fisher's method)

3.9% w/w; calculated for C<sub>23</sub>H<sub>15</sub>N<sub>3</sub>O· $\frac{3}{4}$ H<sub>2</sub>O 3.72% w/w

## EXAMPLE 6

### Production of Anhydrous Crystals of 3-(2-cyanophenyl)-5-(2-pyridyl)-1-phenyl-1,2-dihydropyridin-2-one (Anhydrous Crystal Form V)

A 500 mL-vessel was charged with 9 g of 3-(2-cyanophenyl)-5-(2-pyridyl)-1-phenyl-1,2-dihydropyridin-2-one (crystal form of the hydrate). Acetone (360 mL) was added to the vessel and the mixture was stirred at heating under reflux at 70° C. in a water bath.

After dissolution, the mixture was filtrated by suction and the filtrate was concentrated at 75° C. under normal pressure to solidify. After the solids were finely ground in a mortar, an acetone-water solution prepared from 216 mL of acetone and 54 mL of water was added to the solids.

The mixture was stirred at heating under reflux (at 75° C. in a water bath). After dissolution, the mixture was additional stirred at heating under reflux for 2 hours and 40 minutes. Subsequently, the temperature of the water bath for the mixture (external temperature) was cooled to room temperature at a cooling rate of 10° C./hour, and it was stirred at room temperature for 16 hours.

The precipitates in the reaction mixture were suction-filtrated and then dried under reduced pressure (an external temperature of 20° C. for 40 minutes and then at 60° C. for 3 hours) to give 7.2 g of the anhydrous crystals of 3-(2-cyanophenyl)-5-(2-pyridyl)-1-phenyl-1,2-dihydropyridin-2-one.

<sup>1</sup>H NMR (400 MHz, DMSO-d<sub>6</sub>): δ 8.61-8.57 (m, 1H), 8.53 (d, 1H), 8.47 (d, 1H), 8.01 (d, 1H), 7.92 (dd, 1H), 7.83 (ddd, 1H), 7.78 (ddd, 1H), 7.72 (dd, 1H), 7.61-7.48 (m, 6H), 7.31-7.28 (m, 1H).

## EXAMPLE 7

### Production of the Anhydrous Crystals of 3-(2-cyanophenyl)-5-(2-pyridyl)-1-phenyl-1,2-dihydropyridin-2-one (Anhydrous Crystal Form I)

A 1-L vessel was charged with 8 g of 3-(2-cyanophenyl)-5-(2-pyridyl)-1-phenyl-1,2-dihydropyridin-2-one (Hydrate). Ethyl acetate (480 mL) was added to the vessel and the mixture was stirred at heating under reflux (in an oil bath) to effect dissolution. Heating was stopped and the stirring was allowed to continue while the vessel was in the oil bath (under gradual cooling). At the point that the internal temperature reached 50.9° C., 0.2 g of seed crystals [3-(2-cyanophenyl)-5-(2-pyridyl)-1-phenyl-1,2-dihydropyridin-2-one (anhydrous crystal)] was added to the mixture. Subsequently, stirring continued until the internal temperature reached 31.3° C. The mixture was stirred for additional 2 hours in an ice bath. The precipitated crystals were collected by filtration and dried under aeration (50° C./18 hours) to give 5.8 g of anhydrous crystals of 3-(2-cyanophenyl)-5-(2-pyridyl)-1-phenyl-1,2-dihydropyridin-2-one.

<sup>1</sup>H NMR (400 MHz, DMSO-d<sub>6</sub>): δ 8.58 (d, 1H), 8.53 (d, 1H), 8.47 (d, 1H), 8.01 (d, 1H), 7.93 (d, 1H), 7.83 (ddd, 1H), 7.78 (d, 1H), 7.72 (d, 1H), 7.61-7.48 (m, 6H), 7.32-7.27 (m, 1H).

#### Experiment Test 1

Physical Stability in Mixing-Operation in the Presence of Water or Water-Ethanol (1:1) Mixed Solution

##### (1) Operation Procedure

About 150 mg of respective crystals are placed in an agate mortar, and the mixing operation is carried out at room temperature for a few minutes while the dropwise addition of water (or water-ethanol (1:1) mixed solution) continues.

Subsequently, respective crystals are dried at about 60° C. for 2 to 3 hours.

##### (2) Results

The powder X-ray diffraction analysis showed that the crystals obtained in Reference Example 1 underwent a change in crystal form during the mixing operation in the presence of water or a water-ethanol (1:1) mixed solution and the same crystal form as that obtained in Example 7 increased in quantity.

The powder X-ray diffraction analysis showed that the respective crystals obtained in Example 5, Example 6 and Example 7 displayed no change in crystal form and they were physically stable in the presence of water or a water-ethanol (1:1) mixed solution.

#### Experiment Test 2

Influence of Change in Temperature and Humidity on Hydrate Crystal

##### (1) Apparatus

Rigaku X-ray DTA system: RINT-2000 manufactured by Rigaku Corporation

##### (2) Operation Procedure

Crystals obtained in Example 5 (Hydrate crystal) were ground in a mortar and then sampled on a 13-mm diameter glass plate. Measurement was carried out under the conditions below.

X-ray in use: CuKα ray

Tube voltage: 40 kV

Tube current: 200 mA

Divergent slit: ½ deg

Receiving slit: 0.3 mm

Scattering slit: ½ deg

Scanning speed: 2°/min

Scanning step: 0.01°

Scanning range (2θ): 5 to 40°

Measurement temperature was changed as follows in turn and powder X-ray diffraction patterns were measured at the respective temperatures in sequence: 30→40→50→60→70→80→100→70→60→50→40 and 30° C.

##### (3) Results 1

FIG. 11 shows powder X-ray diffraction patterns of Hydrate at respective temperatures above. The changes in the powder X-ray diffraction patterns revealed that the crystals of Example 5 (Hydrate crystal) transform to the same crystals as crystals of Example E1 (Anhydrous crystal form III) at about 60° C. or more, and return again to Hydrate when the temperature was lowered.

##### (4) Results 2

Measurement humidity was changed as follows in turn and powder X-ray diffraction patterns were measured at the respective humidities in sequence:

4→5→10→15→20→50→90→50→15→5% RH (relative humidity).

FIG. 12 shows powder X-ray diffraction patterns of Hydrate crystal at respective humidities above. From changes

in the powder X-ray diffraction patterns, reversible patterns of Hydrate crystal and Anhydrous crystal form III were observed under the humidity of more and less of about 10% RH. It is confirmed that the crystals of Example 5 (Hydrate crystal) change to Anhydrous crystal form III under the humidity of about 10% RH or less, and stay Hydrate crystal under the humidity of about 10% RH or more.

These experiments about the influence of changes in temperature and humidity on Hydrate crystal, and Example 4 revealed that the state of the precipitates before the air drying was the same crystals as those of Example E1 (Anhydrous crystal form III) or a mixture of Anhydrous crystal form III and Hydrate crystal, which is a useful intermediate for the production of Hydrate crystal.

#### Experiment Test 3

Minimum Ignition Energy and Lower Explosion Concentration Limit

##### (1) Operation Procedure

An appropriate amount of Hydrate crystal corresponding to a concentration was uniformly put on sample dish of blown-up type dust explosion test apparatus. 50 kPa of air was compressed in a 1.3 L pressure tank, and the air was introduced into a glass cylinder by opening of a solenoid operated valve to form dust clouds. A discharge electrode was supplied with energy after 0.1 seconds after the opening of the solenoid operated valve. The criterion of ignition is an arrival of flame at an ignition mark set 100 mm above the discharge electrode.

(2) Measurement Conditions for Lower Explosion Concentration Limit

Temperature of the measurement room: 24° C.

Humidity: 49%

Popping pressure of compressed air: 50 kPa

Ignition start time: 0.1 sec

Repetition of ignition test: 5 times

Ignition discharge energy: 10 J

(3) Measurement Conditions for Minimum Ignition Energy

Temperature of the measurement room: 24° C.

Humidity: 49%

Popping pressure of compressed air: 50 kPa

Ignition start time: 0.1 sec

Repetition of ignition test: 10 times

##### (4) Apparatus

Blown-up type dust explosion test apparatus (Environmental Technology Co., Ltd. DES -10)

##### (5) Results

Lower Explosion Concentration Limit: 160-170 g/m<sup>3</sup>

Minimum Ignition Energy: 50-100 mJ

Dust concentration: 1250 g/m<sup>3</sup>

#### Experiment Test 4

##### Chargeability

##### (1) Operation Procedure

About 1 g of respective compounds is weighed into a weighing bottle (diameter of 35 mm). A stifling bar [fluoroplastic (tetrafluoroethylene resin) coating; 20 mm] is placed in the bottle and after the lid is in place, the powders are stirred for 30 minutes. The lid is opened at the same time that stirring is stopped; and the static potential of the powder is measured using a static potential measuring instrument.

##### (2) Apparatus

STATIRON-DZ3 manufactured by Shishido Electrostatic, Ltd.

##### (3) Results

Crystals of Reference Example 1: 70-100 V

Crystals of Example 5: 0 V

#### Experiment Test 5

Measurement of Infrared Spectra

35

## (1) Operation Procedure

Infrared spectrum of the crystals obtained in Example 5 was measured under the measurement conditions described below according to the potassium bromide disk method for infrared spectrum measurement as described in General Tests in the Japanese Pharmacopoeia.

## (2) Apparatus

FT/IR-620 manufactured by JASCO Corporation

## (3) Conditions for Measurements

Measuring range: 4000-400  $\text{cm}^{-1}$

Resolution: 4  $\text{cm}^{-1}$

Integration number: 36

Scanning speed: 2 mm/sec

## (4) Results

FIG. 1 shows an infrared spectrum of the crystals obtained in Example 5 (KBr method) and FIG. 2 shows an infrared spectrum of the crystals obtained in Example 6 (KBr method).

Table 2 shows the wavenumbers ( $\text{cm}^{-1}$ ) and transmittances (%) of the absorption peaks for the crystals obtained in Example 5.

Table 3 shows the wavenumbers ( $\text{cm}^{-1}$ ) and transmittances (%) of the absorption peaks for the crystals obtained in Example 6.

TABLE 2

PEAK NUMBER	WAVENUMBER ( $\text{cm}^{-1}$ )	% T
1	3406.64	46.3391
2	2217.74	56.7481
3	1661.37	6.6427
4	1619.91	40.6542
5	1588.09	25.9365
6	1566.88	40.1575
7	1550.49	43.7805
8	1482.03	35.2300
9	1434.78	48.9801
10	1369.21	56.9666
11	1318.11	46.9775
12	1282.43	47.1779
13	1249.65	56.0844
14	1157.08	60.4468
15	1099.23	60.6605
16	899.63	68.4385
17	879.38	62.5325
18	784.89	24.6745
19	751.14	35.0064
20	730.89	57.4603
21	697.14	39.9667
22	606.50	56.3319
23	557.33	62.4188
24	505.26	58.3988

TABLE 3

PEAK NUMBER	WAVENUMBER ( $\text{cm}^{-1}$ )	% T
1	3447.13	50.2836
2	3066.26	49.6530
3	2221.59	49.7414
4	1657.52	1.0035
5	1618.95	24.2890
6	1585.20	11.2133
7	1565.92	26.5039
8	1549.52	26.7864
9	1476.24	17.5093
10	1434.78	36.9734
11	1368.25	45.1127
12	1318.11	29.2805
13	1266.04	36.7269
14	1247.72	39.2277
15	1157.08	52.6189

36

TABLE 3-continued

PEAK NUMBER	WAVENUMBER ( $\text{cm}^{-1}$ )	% T
16	1135.87	64.5032
17	1097.30	55.2787
18	933.38	66.5017
19	896.74	61.6446
20	879.38	55.0085
21	785.85	14.4121
22	753.07	24.8345
23	729.92	54.7682
24	694.25	30.5167
25	630.61	61.1148
26	603.61	46.5267
27	556.36	62.4971
28	501.40	51.0929
29	443.55	67.0781

## Experiment Test 6

## Measurement of Powder X-Ray Diffraction Pattern

## Operation Procedure

Powder X-ray diffraction patterns of the crystals obtained in the respective Examples were measured under the measurement conditions described below according to the powder X-ray diffraction measurement method as described in General Tests in the Japanese Pharmacopoeia.

## (2) Apparatus

Rigaku X-ray DTA system: RINT-2000 manufactured by Rigaku Corporation

## (3) Operation Procedure

A sample was ground in a mortar and then sampled on a 13-mm diameter glass plate. Measurement was carried out under the conditions below.

X-ray in use:  $\text{CuK}\alpha$  ray

Tube voltage: 40 kV

Tube current: 200 mA

Divergent slit:  $\frac{1}{2}$  deg

Receiving slit: 0.3 mm

Scattering slit:  $\frac{1}{2}$  deg

Scanning speed: 1°/min

Scanning step: 0.01°

Scanning range (2 $\theta$ ): 5 to 40°

## (4) Results

FIG. 3 shows a powder X-ray diffraction pattern of the crystals obtained in Reference Example 1, FIG. 4 shows a powder X-ray diffraction pattern of the crystals obtained in Example 5, FIG. 5 shows a powder X-ray diffraction pattern of the crystals obtained in Example 6, and FIG. 6 shows a powder X-ray diffraction pattern of the crystals obtained in Example 7.

Table 4 shows the peaks and their intensities at diffraction angles (2 $\theta$ ) for the crystals obtained in Reference Example 1, Table 5 shows the peaks and their intensities at diffraction angles (2 $\theta$ ) for the crystals obtained in Example 5, Table 6 shows the peaks and their intensities at diffraction angles (2 $\theta$ ) for the crystals obtained in Example 6, and Table 7 shows the peaks and their intensities at diffraction angles (2 $\theta$ ) for the crystals obtained in Example 7.

Based on FIG. 4 and Table 5 that represent the powder X-ray diffraction pattern of the crystals obtained in Example 5, it can be found that the powder X-ray diffraction pattern of the crystals obtained in Example 5 provides a characteristic peak having a diffraction angle (2 $\theta$ ) of about 12.5°.

This suggests that the crystals obtained in Reference Example 1 do not contain the same crystal form as do the crystals obtained in Example 5, since FIG. 3 and Table 4 that represent the powder X-ray diffraction pattern of the crystals

37

obtained in Reference Example 1 does not provide the peak having a diffraction angle ( $2\theta$ ) of about  $12.5^\circ$ .

## EXAMPLE E1

## Anhydrous Form III

With respect to the crystals of 3-(2-cyanophenyl)-5-(2-pyridyl)-1-phenyl-1,2-dihydropyridin-2-one hydrate, the powder X-ray diffraction pattern was measured under the conditions similar to those described above. However, the measurement was conducted at a scanning speed of  $2^\circ/\text{min}$  under heating conditions in the vicinity of  $110^\circ\text{C}$ .

FIG. 7 shows the powder X-ray diffraction pattern, and Table 8 shows the peaks and the intensity of diffraction angles ( $2\theta \pm 0.2^\circ$ ) for the crystals.

TABLE 4

PEAK NUMBER	$2\theta$	HALF WIDTH	d-VALUE	INTENSITY	RELATIVE INTENSITY
1	9.010	0.588	9.8067	13370	100
2	15.850	0.682	5.5867	10137	76
3	24.390	0.847	3.6465	10672	80

TABLE 5

PEAK NUMBER	$2\theta$	HALF WIDTH	d-VALUE	INTENSITY	RELATIVE INTENSITY
1	7.780	0.259	11.3542	16328	100
2	8.700	0.247	10.1555	5298	32
3	9.520	0.282	9.2825	11203	69
4	12.450	0.365	7.1037	4845	30
5	14.590	0.282	6.0663	2872	18
6	15.240	0.329	5.8090	7037	43
7	15.600	0.188	5.6757	2162	13
8	16.180	0.282	5.4735	2358	14
9	17.540	0.341	5.0521	6268	38
10	19.980	0.318	4.4403	3823	23
11	21.040	0.247	4.2189	2430	15
12	21.420	0.271	4.1449	2708	17
13	23.260	0.282	3.8210	3838	24
14	24.190	0.294	3.6762	12030	74
15	25.440	0.282	3.4983	5205	32
16	25.820	0.188	3.4477	2953	18
17	26.430	0.294	3.3695	4488	27
18	27.310	0.259	3.2629	2295	14
19	27.860	0.341	3.1997	3567	22
20	28.870	0.294	3.0900	2645	16

TABLE 6

PEAK NUMBER	$2\theta$	HALF WIDTH	d-VALUE	INTENSITY	RELATIVE INTENSITY
1	8.000	0.259	11.0424	18318	76
2	9.160	0.259	9.6465	24045	100
3	10.000	0.294	8.8380	3002	12
4	12.870	0.259	6.8728	2838	12
5	15.140	0.247	5.8471	5210	22
6	15.950	0.282	5.5519	17707	74
7	16.720	0.306	5.2979	8272	34
8	18.370	0.271	4.8256	5460	23
9	18.820	0.235	4.7113	4523	19
10	20.650	0.259	4.2977	5878	24
11	22.150	0.318	4.0099	3967	16
12	23.100	0.224	3.8471	5378	22
13	24.320	0.365	3.6568	5333	22
14	24.930	0.306	3.5687	19343	80
15	26.280	0.271	3.3884	3737	16
16	26.860	0.282	3.3165	5103	21

38

TABLE 6-continued

PEAK NUMBER	$2\theta$	HALF WIDTH	d-VALUE	INTENSITY	RELATIVE INTENSITY
17	28.010	0.306	3.1829	2182	9
18	28.660	0.294	3.1122	5767	24
19	30.410	0.259	2.9369	3333	14
20	30.900	0.282	2.8915	2363	10

TABLE 7

PEAK NUMBER	$2\theta$	HALF WIDTH	d-VALUE	INTENSITY	RELATIVE INTENSITY
1	7.780	0.235	11.3542	20997	100
2	9.450	0.259	9.3511	7993	38
3	10.270	0.247	8.6062	6090	29
4	14.240	0.247	6.2146	3092	15
5	15.060	0.353	5.8780	6283	30
6	16.070	0.271	5.5107	16920	81
7	19.080	0.294	4.6476	7162	34
8	20.450	0.259	4.3393	3658	17
9	22.440	0.259	3.9588	3212	15
10	23.300	0.176	3.8145	5342	25
11	24.150	0.094	3.6822	7478	36
12	27.310	0.106	3.2629	4648	22

TABLE 8

PEAK NUMBER	$2\theta$	HALF WIDTH	d-VALUE	INTENSITY	RELATIVE INTENSITY
1	5.700	0.235	15.4919	557	8
2	7.660	0.224	11.5318	6780	100
3	8.650	0.247	10.2141	1813	27
4	9.490	0.259	9.3118	3490	51
5	11.640	0.259	7.5962	1603	24
6	12.350	0.282	7.1610	1817	27
7	13.690	0.271	6.4630	1057	16
8	14.550	0.200	6.0828	1247	18
9	15.240	0.271	5.8090	1230	18
10	15.530	0.118	5.7011	717	11
11	16.110	0.141	5.4972	650	10
12	17.590	0.306	5.0378	1237	18
13	19.940	0.141	4.4491	417	6
14	21.450	0.141	4.1392	540	8
15	22.300	0.271	3.9833	750	11
16	23.700	0.247	3.7511	2560	38
17	24.960	0.235	3.5645	1783	26
18	25.420	0.106	3.5010	1173	17
19	26.150	0.259	3.4049	720	11
20	27.030	*****	3.2960	703	10
21	28.500	0.200	3.1293	670	10

## Experiment Test 7

Measurement of  $^{13}\text{C}$  Solid State NMR Spectrum

## (1) Operation Procedure

$^{13}\text{C}$  Solid State NMR spectra were measured for the crystals obtained in Examples 5, 6 and 7 under the following conditions.

Measurement temperature: room temperature ( $\sim 22^\circ\text{C}$ .)

Standard compound: carbonyl carbon of glycine (external standard: 176.03 ppm)

Measurement nucleus:  $^{13}\text{C}$  (100.6248425 MHz)

Pulse-repetition time: 50 sec for Examples 6 and 7 5 sec for Example B1

Pulse mode: CP/TOSS measurement

## (2) Results

FIG. 8 shows a  $^{13}\text{C}$  Solid State NMR spectrum of the crystals obtained in Example 5, and the chemical shifts are summarized in Table 9. FIG. 9 shows a  $^{13}\text{C}$  Solid State NMR spectrum of the crystals obtained in Example 7, and the

chemical shifts are summarized in Table 10. FIG. 10 shows a  $^{13}\text{C}$  Solid State NMR spectrum of the crystals obtained in Example 6, and the chemical shifts are summarized in Table 11.

TABLE 9

Chemical Shift(ppm)
159.8
150.8
146.7
139.4
136.6
134.9
132.7
129.0
127.8
126.5
125.8
123.3
120.8
118.4
114.6
108.8

TABLE 10

Chemical Shift(ppm)
160.0
152.8
152.0
149.0
140.5
137.0
134.7
133.4
131.0
129.5
128.2
127.0
125.6
120.3
118.2
114.6
108.2

TABLE 11

Chemical Shift(ppm)
159.7
151.2
145.9
139.9
137.7
134.9
132.8
131.7
130.1
129.0
128.2
127.3
126.8
125.8
121.1
120.6
119.1
114.2
109.4

As explained in detail herein above, according to the present invention, a compound represented by formula (III) may be industrially produced in good yield and high purity by reacting a compound represented by formula (I) with a com-

pound represented by formula (II) in the presence of a palladium compound, a copper compound, a phosphorus compound and a base.

The crystal forms of the present invention have preferable properties and are suitable for use as an active ingredient of therapeutic or prophylactic agents for neurodegenerative diseases or the like.

What is claimed is:

1. An anhydrous crystal of 3-(2-cyanophenyl)-5-(2-pyridyl)-1-phenyl-1,2-dihydropyridin-2-one having a diffraction peak at a diffraction angle ( $2\theta \pm 0.2^\circ$ ) of  $10.3^\circ$  in a powder X-ray diffraction.

2. The crystal according to claim 1 further having a diffraction peak at a diffraction angle ( $2\theta \pm 0.2^\circ$ ) of  $19.1^\circ$  in a powder X-ray diffraction.

3. An anhydrous crystal of 3-(2-cyanophenyl)-5-(2-pyridyl)-1-phenyl-1,2-dihydropyridin-2-one having peaks at chemical shifts of around 149.0 ppm and around 125.6 ppm in a  $^{13}\text{C}$  Solid State Nuclear Magnetic Resonance spectrum.

4. An anhydrous crystal of 3-(2-cyanophenyl)-5-(2-pyridyl)-1-phenyl-1,2-dihydropyridin-2-one having a diffraction peak at a diffraction angle ( $2\theta \pm 0.2^\circ$ ) of  $16.7^\circ$  in a powder X-ray diffraction.

5. The crystal according to claim 4 further having diffraction peaks at diffraction angles ( $2\theta \pm 0.2^\circ$ ) of  $12.9^\circ$  and  $24.9^\circ$  in a powder X-ray diffraction.

6. An anhydrous crystal of 3-(2-cyanophenyl)-5-(2-pyridyl)-1-phenyl-1,2-dihydropyridin-2-one having an absorption peak at a wavenumber of  $1658 \pm 1 \text{ cm}^{-1}$  in an infrared absorption spectrum (KBr method).

7. The crystal according to claim 6 further having an absorption peak at a wavenumber of  $501 \pm 1 \text{ cm}^{-1}$  in an infrared absorption spectrum (KBr method).

8. An anhydrous crystal of 3-(2-cyanophenyl)-5-(2-pyridyl)-1-phenyl-1,2-dihydropyridin-2-one having peaks at chemical shifts of around 145.9 ppm and around 137.7 ppm in a  $^{13}\text{C}$  Solid State Nuclear Magnetic Resonance spectrum.

9. An anhydrous crystal of 3-(2-cyanophenyl)-5-(2-pyridyl)-1-phenyl-1,2-dihydropyridin-2-one having diffraction peaks at diffraction angles ( $2\theta \pm 0.2^\circ$ ) of  $23.7^\circ$  and  $25.0^\circ$  in a powder X-ray diffraction.

10. The crystal according to claim 6 further having diffraction peaks at diffraction angles ( $2\theta \pm 0.2^\circ$ ) of  $5.7^\circ$  and  $9.5^\circ$  in a powder X-ray diffraction.

11. An anhydrous crystal of 3-(2-cyanophenyl)-5-(2-pyridyl)-1-phenyl-1,2-dihydropyridin-2-one having a diffraction peak at a diffraction angle ( $2\theta \pm 0.2^\circ$ ) of  $10.3^\circ$ ,  $14.2^\circ$  and  $19.1^\circ$  in a powder X-ray diffraction.

12. An anhydrous crystal of 3-(2-cyanophenyl)-5-(2-pyridyl)-1-phenyl-1,2-dihydropyridin-2-one having a diffraction peak at a diffraction angle ( $2\theta \pm 0.2^\circ$ ) of  $7.8^\circ$ ,  $9.5^\circ$ ,  $10.3^\circ$ ,  $14.2^\circ$ ,  $15.1^\circ$ ,  $16.1^\circ$ ,  $19.1^\circ$  and  $20.5^\circ$  in a powder X-ray diffraction.

13. An anhydrous crystal of 3-(2-cyanophenyl)-5-(2-pyridyl)-1-phenyl-1,2-dihydropyridin-2-one having peaks at chemical shifts of around 160.0 ppm, 152.8 ppm, 152.0 ppm, 149.0 ppm, 140.5 ppm, 137.0 ppm, 134.7 ppm, 133.4 ppm, 131.0 ppm, 129.5 ppm, 128.2 ppm, 127.0 ppm, 125.6 ppm, 120.3 ppm, 118.2 ppm, 114.6 ppm and 108.2 ppm in a  $^{13}\text{C}$  Solid State Nuclear Magnetic Resonance spectrum.

14. An anhydrous crystal of 3-(2-cyanophenyl)-5-(2-pyridyl)-1-phenyl-1,2-dihydropyridin-2-one having a diffraction peak at a diffraction angle ( $2\theta \pm 0.2^\circ$ ) of  $12.9^\circ$ ,  $16.7^\circ$  and  $24.9^\circ$  in a powder X-ray diffraction.

15. An anhydrous crystal of 3-(2-cyanophenyl)-5-(2-pyridyl)-1-phenyl-1,2-dihydropyridin-2-one having a diffraction

## 41

tion peak at a diffraction angle ( $2\theta \pm 0.2^\circ$ ) of  $8.0^\circ$ ,  $9.2^\circ$ ,  $10.0^\circ$ ,  $12.9^\circ$ ,  $16.7^\circ$ ,  $18.4^\circ$ ,  $18.8^\circ$ ,  $20.7^\circ$  and  $24.9^\circ$  in a powder X-ray diffraction.

16. An anhydrous crystal of 3-(2-cyanophenyl)-5-(2-pyridyl)-1-phenyl-1,2-dihydropyridin-2-one having peaks at chemical shifts of around 159.7 ppm, 151.2 ppm, 145.9 ppm, 139.9 ppm, 137.7 ppm, 134.9 ppm, 132.8 ppm, 131.7 ppm, 130.1 ppm, 129.0 ppm, 128.2 ppm, 127.3 ppm, 126.8 ppm, 125.8 ppm, 121.1 ppm, 120.6 ppm, 119.1 ppm, 114.2 ppm and 109.4 ppm in a  $^{13}\text{C}$  Solid State Nuclear Magnetic Resonance spectrum.

17. An anhydrous crystal of 3-(2-cyanophenyl)-5-(2-pyridyl)-1-phenyl-1,2-dihydropyridin-2-one having an absorption peak at a wavenumber of  $3447.1 \pm 0.1 \text{ cm}^{-1}$ ,  $3066.3 \pm 0.1 \text{ cm}^{-1}$ ,  $2221.6 \pm 0.1 \text{ cm}^{-1}$ ,  $1657.5 \pm 0.1 \text{ cm}^{-1}$ ,  $1585.2 \pm 0.1 \text{ cm}^{-1}$ ,  $1476.2 \pm 0.1 \text{ cm}^{-1}$  and  $785.9 \pm 0.1 \text{ cm}^{-1}$  in an infrared absorption spectrum (KBr method).

18. An anhydrous crystal of 3-(2-cyanophenyl)-5-(2-pyridyl)-1-phenyl-1,2-dihydropyridin-2-one having an absorption peak at a wavenumber of  $3447.1 \pm 0.1 \text{ cm}^{-1}$ ,  $3066.3 \pm 0.1$

## 42

$\text{cm}^{-1}$ ,  $2221.6 \pm 0.1 \text{ cm}^{-1}$ ,  $1657.5 \pm 0.1 \text{ cm}^{-1}$ ,  $1619.0 \pm 0.1 \text{ cm}^{-1}$ ,  $1585.2 \pm 0.1 \text{ cm}^{-1}$ ,  $1565.9 \pm 0.1 \text{ cm}^{-1}$ ,  $1549.5 \pm 0.1 \text{ cm}^{-1}$ ,  $1476.2 \pm 0.1 \text{ cm}^{-1}$ ,  $1434.8 \pm 0.1 \text{ cm}^{-1}$ ,  $1368.3 \pm 0.1 \text{ cm}^{-1}$ ,  $1318.1 \pm 0.1 \text{ cm}^{-1}$ ,  $1266.0 \pm 0.1 \text{ cm}^{-1}$ ,  $1247.7 \pm 0.1 \text{ cm}^{-1}$ ,  $1157.1 \pm 0.1 \text{ cm}^{-1}$ ,  $1135.9 \pm 0.1 \text{ cm}^{-1}$ ,  $1097.3 \pm 0.1 \text{ cm}^{-1}$ ,  $933.4 \pm 0.1 \text{ cm}^{-1}$ ,  $896.7 \pm 0.1 \text{ cm}^{-1}$ ,  $879.4 \pm 0.1 \text{ cm}^{-1}$ ,  $785.9 \pm 0.1 \text{ cm}^{-1}$ ,  $753.1 \pm 0.1 \text{ cm}^{-1}$ ,  $729.9 \pm 0.1 \text{ cm}^{-1}$ ,  $694.3 \pm 0.1 \text{ cm}^{-1}$ ,  $630.6 \pm 0.1 \text{ cm}^{-1}$ ,  $603.6 \pm 0.1 \text{ cm}^{-1}$ ,  $556.4 \pm 0.1 \text{ cm}^{-1}$ ,  $501.4 \pm 0.1 \text{ cm}^{-1}$  and  $443.6 \pm 0.1 \text{ cm}^{-1}$  in an infrared absorption spectrum (KBr method).

19. An anhydrous crystal of 3-(2-cyanophenyl)-5-(2-pyridyl)-1-phenyl-1,2-dihydropyridin-2-one having a diffraction peak at a diffraction angle ( $2\theta \pm 0.2^\circ$ ) of  $5.7^\circ$ ,  $9.5^\circ$ ,  $23.7^\circ$  and  $25.0^\circ$  in a powder X-ray diffraction.

20. An anhydrous crystal of 3-(2-cyanophenyl)-5-(2-pyridyl)-1-phenyl-1,2-dihydropyridin-2-one having a diffraction peak at a diffraction angle ( $2\theta \pm 0.2^\circ$ ) of  $5.7^\circ$ ,  $7.7^\circ$ ,  $8.7^\circ$ ,  $9.5^\circ$ ,  $11.6^\circ$ ,  $12.4^\circ$ ,  $17.6^\circ$ ,  $23.7^\circ$  and  $25.0^\circ$  in a powder X-ray diffraction.

\* \* \* \* \*